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(54) Title: PROCESS FOR PREPARATION OF ALOE PRODUCTS (57) Abstract <p>A process is described for extracting a pharmaceutically active polysaccharidic substance from the aloe plant. The pharmaceutically active polysaccharidic substance and its characteristic properties are described.</p>		

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PROCESS FOR PREPARATION OF ALOE PRODUCTS

RELATED APPLICATIONS

The present application is a continuation-in-part of U.S. application Serial No. 869,261, the entire contents and disclosure of which are hereby specifically incorporated by reference. Said U.S. application Serial No. 869,261 corresponds to International Application No. PCT/US86/01335 filed June 20, 1986 and published under International Publication No. WO 87/00052 on January 15, 1987, the entire contents and disclosure of which are also hereby specifically incorporated by reference. Said U.S. application Serial No. 869,261 is a continuation-in-part of U.S. application Serial No. 810,025 filed December 17, 1985 (now abandoned), which was a continuation-in-part of U.S. application Serial No. 754,859 filed July 12, 1985 (now abandoned), which was a continuation-in-part of U.S. application Serial No. 750,321 filed June 28, 1985 (now abandoned), which was a continuation-in-part of U.S. application Serial No. 649,967 filed September 12, 1984 (now abandoned), which was a continuation of U.S. application Serial No. 375,720 filed May 7, 1982 (now abandoned). Said application Serial No. 869,261 being entitled "Processes For Preparation of Aloe Products, Products Produced Thereby and Compositions Thereof". Said application Serial No. 810,025 being entitled "Processes for Preparation of Aloe Products and Products Produced Thereby". Said applications Serial Nos. 754,859; 750,321; 649,967; and 375,720 being entitled "Process for Preparation of Aloe Vera Products".

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BACKGROUND OF THE INVENTION

1. Field of the Invention

This invention pertains to the field of processing aloe plants and removing portions of said plant for processing same into compositions for topical and internal applications, and compositions of matter comprising said portions of aloe.

2. Description of the Prior Art, and Other Information

Aloe vera is not a cactus plant as widely believed, but rather a member of the lily family. There are about 360 species of aloe plants known. Harding, Aloes of the World: A Checklist, index and code, Excelsa 9: 57 - 94 (1979). They seem to thrive in hot arid areas and are widely scattered from the Mediterranean Sea, Middle East, Africa, China, Japan, Mexico and the southern U.S.A. A few of the important species used for their medicinal properties are Aloe barbadensis Miller (aloe vera), A. arborescens, A. plicatilis, A. vahombe, A. saponaria, A. africana, A. ferox and Aloe perryi. Reynolds, Aloes of Tropical Africa and Madagascar, The Trustees, The Aloe Book Fund, Mbabane Swaziland. However, A. barbadensis Miller is generally recognized as the "true aloe" because of its wide use and reportedly most effective healing power, although in Japan, A. arborescens Miller traditionally has been used as a folk remedy for various ailments ranging from gastrointestinal disorders to athlete's foot.

Aloe vera is a perennial plant with turgid green leaves joined at the stem in a rosette pattern. The leaves of a mature plant may be more than 25 inches long with saw-like spikes along their margins.

Slicing the leaf transversely as shown in Figure 1 and 2 reveals the outer walls of the epidermis 3 covered with thick cuticles. Beneath the epidermis 3 is the mesophyll which is differentiated into chlorenchyma cells and thinner walled cells known as parenchyma. The parenchyma cells harbor a transparent mucilaginous jelly 1. The vascular bundles 2 with inner bundle sheath cells contain the yellow sap having

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laxative properties and are sandwiched between the two major cells. Needle shaped crystals of calcium oxalate, produced as a metabolic byproduct in plant cells, are found mostly at the central portion of the leaf.

Aloe vera contains two major liquid sources, a yellow latex (exudates) and the clear gel (mucilage). The dried exudate of Aloe barbadensis Miller leaves is referred to as Aloe. The commercial name is Curacao aloe. It is composed mainly of aloin, aloe-emodin and phenols. Bruce, South African Medical Journal, 41: 984 (1967); Morrow et al., Arch. Dermatology, 116: 1064-1065 (1980); Salek et al., Corrosion Prevention & Control, 9 - 10 (1983); Mapp et al., Planta Medica, 18: 361 - 365 (1970); Ranwald, Arch. Pharmacology, 315: 477 - 478 (1982). A number of phenolics, including anthraquinones and their glycosides, are known to be pharmaceutically active. Bruce, Excelsa 5: 57 - 68 (1975); Suga et al., Cosmetic and Toiletries, 98: 105 - 108 (1983).

The mucilaginous jelly from the parenchyma cells of the plant is referred to as aloe vera gel. There are generally no anthraquinones to decompose and cause discoloration of the gel unless the gel is contaminated by an improper processing technique.

Aloe vera gel is about 98.5% by weight water. More than 60% of the total solid is made up of polysaccharides of carbohydrate origin. Organic acids and inorganic compounds, especially calcium oxalate, account for the remainder of the solid.

Whole leaves, exudates and fresh gels of aloe plants have been used for a variety of human afflictions. Evidence of their use as a medicinal remedy can be traced to the Egyptians of 400 BC. Aloe vera was also used to embalm the dead as well as to protect the embalmers from the death-causing agent. Other early civilizations used aloe vera for skin care, relieving insect stings and bites, treating scratches, wound healing, hair loss, as a purgative and for ulcerated skin. It was the traditional medicine of many

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cultures as an anthelmintic, cathartic, stomachic, and was used inter alia for leprosy, burns and allergic conditions. Cole et al. Archives of Dermatology and Syphilology, 47: 250 (1943); Chopra et al., Glossary of Indian Medicinal Plants, Council of Scientific and Industrial Research, New Delhi; Ship, Journal of American Medical Association, 238: 1770 - 1772 (1977); Morton, Atlas of Medicinal Plants of Middle American Bahamas to Yucatan, Charles C. Thomas ED., 78 - 80 (1981); Diez-Martinez, La Zabala, Comunicado NO. 46 Sobre Recursos Bioticos Potenciales del Pais, INIREB, Mexico (1981); Dastur, Medicinal Plants of India and Pakistan; D.B. Taraporevala Sons & Co., Private Ltd., Bombay 16 - 17 (1962).

Aloe vera has enjoyed a long history of lay acceptance as possessing "curative" or "healing qualities". Over the last few years, numerous books and articles meeting scientific standards have been written on Aloe vera. Organizations such as the Aloe Vera Council and recognized medical institutions through publications and case-histories of physicians, veterinarians and other scientists have given credence to the "aloe phenomenon". Aloe vera has been featured extensively in the area of dermatology, especially for treating radiation-caused skin conditions. Mackee, X-Rays and Radium in the Treatment of Diseases of the Skin, 3rd Ed., Lea and Febiger, Philadelphia, 319 - 320 (1938); Rovalti et al., Industrial Medicine and Surgery, 28: 364 - 368 (1959); Zawahry et al., Quotations From Medical Journals on Aloe Research, Ed. Max B. Skousen, Aloe Vera Research Institute, Cypress, California 18 - 23 (1977); Cera et al., Journal of the American Animal Hospital Association, 18: 633 - 638 (1982). The body of scientific literature documenting medical applications in digestive problems, as a virucidal, bactericidal, and fungicidal agent and in gynecological conditions is extensive and has been adequately reviewed by Grindley and Reynolds (Journal of Ethnopharmacology, 16: 117 - 151 (1986)).

The importance of chemicals found in aloes is indicated by the fact that they have been listed in every known

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national pharmacopeia. U.S. Pharmacopeia, 20th Revision, The National Formulary, 14th Edition, United States Pharmacopeial Convention, Inc., Rockville, Maryland, July 1, 1980. However, the U.S. Pharmacopeia describes the yellow sap drug portion of aloes but not the mucilage. The fresh unpreserved gel is about 98.5-99.2 percent water. The total solid that remains after the water has been removed ranges from 0.8 to 1.5 percent. The major constituents of the solid comprise mucilage, sugars, fiber, proteins, ash, fats, aloin and resin. Robson et al., Journal of Burn Care Rehabilitation, 3: 157 - 163 (1982). Reports of compositions which include enzymes, organic acids, inorganic salts, amino acids and alkaloids have been noted. Rowe et al., Journal of the American Pharmaceutical Association, 30: 262 - 266 (1941); Roboz et al., Journal of the American Chemical Society, 70: 3248 - 3249 (1948); Waller et al., Proceedings of Oklahoma Academy of Science, 58: 69 - 76 (1978). Depending on how the leaves are processed, mucilage and sugars are the major components of the dehydrated gel. The sugars found are galactose, glucose, mannose, rhamnose, xylose and uronic acids. Although conflicting reports have been observed, the mucilage is mainly composed of mannan or glucomannan. Eberendu et al., The Chemical Characterization of Carrisyn™ (in preparation); Mandal et al., Carbohydrate Research, 87: 249 - 256 (1980b); Roboz et al., Journal of the American Chemical Society, 70: 3248 - 3249 (1948); Gowda et al., Carbohydrate Research, 72: 201 - 205 (1979); Segal et al., Lloydia, 31: 423 (1968).

Presently, the controversy over the identity of the active substance(s) in aloe vera has not been settled. It is therefore important to clearly distinguish between the components present in the gel and those found in the exudates. A larger part of the gel is a mucilage of mainly polysaccharide nature with minor amounts of various other compounds. It is conceivable that there may be some synergistic action between the polysaccharide base and other components in some of the activities observed. Leung, Excelsa 8: 65 - 68 (1978); Henry,

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Cosmetic and Toiletries, 94: 42 - 50 (1979). For example, several workers report that the effective components for wound healing may be traumatic acid (Freytag, Pharmazie, 9: 705 (1954)) and a kind of polysaccharide. Kawashima et al., Chemical Abstracts, 93: 13075 (1979). Mackee, supra, noted that the gel not the rind or the exudate, was responsible for the beneficial effects in the treatment of radiation burns, and he stressed the importance of using fresh leaves for effective treatment. Polysaccharides degrade with time and certain molecular weight sizes may be necessary to elicit specified pharmacological response. Goto et al., Gann, 63: 371 - 374 (1972).

There are many examples in the literature of polysaccharides demonstrating pharmacological and physiological activities without known synergistic help from other components. Ohno et al., Chem. Pharm. Bull., 33: 2564 - 2568 (1985); Leibovici et al., Chem. Biol. Interactions, 60: 191 - 200 (1986); Ukai et al., Chem. Pharm. Bull., 31: 741 - 744 (1983); Leibovici et al., Anticancer Research, 5: 553 - 558 (1985). Hyperlipidemia in the general population and especially in familial hypercholesterolemia is associated with coronary heart disease and death. In countries where dietary fiber intake is high, atherosclerosis appears to be uncommon. Trowell et al., Editors, Refined Carbohydrate Foods and Disease, London, Academic Press, 207 (1975). Pectin and guar are reported to lower cholesterol in normal and hyperlipidemic patients. Kay et al., American Journal of Clinical Nutrition, 30: 171 - 175 (1977). Locust bean gum, a polysaccharide composed of mannose and galactose, decreased the lipoprotein cholesterol levels in cases of normal and familial hypercholesterolemic subjects. Zavoral et al., American Journal of Clinical Nutrition, 38: 285 - 294 (1983). Addition of guar gum to carbohydrate meals decreased the postprandial rise of glucose in both normal and diabetic subjects. Jenkins et al., Lancet, 2: 779 - 780 (1977). Kuhl et al., (Diabetes Care), 6 (2): 152 - 154 (1983)) demonstrated that guar gum

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exhibited glycemic control of pregnant insulin dependent diabetic patients.

Anti-tumor activity of polysaccharides is widely reported. Polysaccharides prepared from Lentinus cyathiformis are known to increase hosts' defense against tumors. Rethy et al., Annals of Immunology Hungary, 21: 285 - 290 (1981). There are several reports of polysaccharides from mushroom, yeast or bacterium extracts eliciting a high degree of host defense activity against viral and tumor infestations. Chihara et al., Nature, 222: 687 (1969); Schwartzman, Proc. Soc. Exper. Biol. Med., 29: 737 (1932); Rethy, X. International Congress of Microbiology; Moscow, 642 (1966). Suzuki et al. (Journal of Pharm. Dyn., 7: 492 - 500 (1984)) also reported anti-tumor activity of a polysaccharide fraction extracted from cultured fruiting bodies of a fungus, Grifola frondosa. The fraction (GF-1) showed equivalent levels of higher inhibiting activity when administered intraperitoneally (IP), intravenously (IV) and intratumorally (IT). However, oral administration (PO) was reported not effective. The GF-1 also exhibited anti-tumor action against the solid form of Meth A fibrosarcoma and MM 46 Carcinoma in mice. Lentinan, which is a 6-branched B-1-3-linked glucan similar to GF-1, was ineffective against Meth A fibrosarcoma. Chihara, The antitumor polysaccharide Lentinan: an overview; "Manipulation of Host Defense Mechanisms"; ed. by Aoki et al., Excerpta Medica, North Holland, 1 - 16 (1981); Sasaki et al., Carbohydrate Research, 47: 99 - 104 (1976). Synthesized branched polysaccharides were reported to demonstrate bioactivities against tumors. Matsuzaki et al., Makromol. Chem., 186: 449 (1985). Matsuzaki et al. (Makromol. Chem., 187: 325 - 331 (1986)) synthesized branched polysaccharides from ivory nut mannan, B-(1-4)-D-mannopyranose and B-(1-4)-linked glucomannan which showed significant activities. A partially acetylated linear B-(1-3)-D-mannan extracted from fruit bodies of Dictyophoria Indusiata Fisch exhibited anti-tumor activity. Hara et al., Carbohydrate Research, 143: 111 (1982). It appears that anti-

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tumor action depends on the kind of polymer main chain and its degree of polymerization since B-(1-3)-glucan type polymers show higher anti-tumor activity than B-(1-4)-glucan and hemi-cellulosic polymers. Matsuzaki et al., Makromol. Chem., 187: 325 - 331 (1986). A carboxymethylated derivative of B-(1-3)-glucan obtained from bacterial culture filtrate caused severe cell loss from established sarcoma 180 tumors within two hours after the injection of the derivative. Baba et al., J. of Immunopharmacology, 8: 569 - 572 (1986). The same author observed a compensatory increase in polymorphonuclear leukocytes due to the injection of the substance. Incidentally bestatin, a dipeptide known to possess immune-modulating and anti-tumor activity (Ishizuka et al., Journal of Antibiotics, 32: 642 - 652 (1980), influenced neither the tumor yield nor the polymorphonuclear leukocyte count. Baba et al., supra.

There are numerous reports on the anti-tumor effect of sulfated polysaccharides, including heparain (Jolles et al., Acta Univ. Int. Cancer, 16: 682 - 685 (1960); Suemasu et al., Gann, 61: 125 - 130 (1970)), sulfated laminaran and dextran. Jolles et al., British Journal of Cancer, 17: 109 - 115 (1963). Yamamoto et al. (Japan Journal of Experimental Medicine, 54: 143 - 151 (1984)) reported enhancement of anti-tumor activity of a fucoidan fraction by further sulfation. The sulfated product demonstrated activity against L-1210 leukemia. The authors postulated that the mechanism of the anti-tumor action might be due partly to inhibition of invasive growth of L-1210 cells resulting from electrostatic repulsion between the tumor cell and mesothelial cells. Yamamoto et al., supra. Polysaccharides with sulfate groups are also reported to be human T cell mitogens and murine polyclonal B cell activators. Sugawara et al., Microbiological Immunology, 28: 831 - 839 (1984). Generally, homopolysaccharides of high molecular weight with sulfate groups possess these properties. Dorries et al., European Journal of Immunology, 4: (19); Sugawara et al., Cell Immunology, 74: 162 - 171 (1982).

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It has been reported that glucan extracted from the yeast Saccharomyces cerevisiae is a modulator of cellular and humoral immunity. Wooles et al., Science, 142: 1078 - 1080 (1963). The polysaccharide also stimulated the proliferation of murine pluripotent hematopoietic stem cells, granulocyte-macrophage colony-forming cells, and cells forming myeloid and erythroid colonies. Pospisil et al., Experientia, 38: 1232 - 1234 (1982); Burgaleta et al., Cancer Research, 37: 1739 - 1742 (1977). Maisin et al. (Radiation Research, 105: 276 - 281 (1986)) also reported an IV administration of a polysaccharide that induced protection of Murine hematopoietic stem cells against x-ray exposure, thereby decreasing the mortality of the mice so exposed.

Seljelid et al., (Experimental Cell Research, 131: 121 (1981)) have observed that insoluble or gelforming glycans activated macrophages in vitro whereas the corresponding soluble glycans did not. They postulated that the way the glycan was presented to the mononuclear phagocyte was decisive for activation. Bogwald et al. (Scand. Journal of Immunology, 20: 355 - 360 (1984)) immobilized glycans which had a stimulatory effect on the macrophages in vitro. This led the authors to believe that the degree of fixed or steric arrangement of the glycan was decisive for the effect on the macrophages in vitro. A purified polysaccharide isolated from Candida albicans induced antibody response in vitro by human peripheral blood lymphocytes. Wirz et al., Clinical Immunology and Immunopathology, 33: 199 - 209 (1984), There were significant differences between the anti-candida antibodies in sera of normal and candida-infected individuals. Wirz et al., supra.

The antiviral activity of polysaccharides and polysaccharides linked to peptides has been observed. Suzuki et al., Journal Antibiotics, 32: 1336 - 1345 (1979). Suzuki et al., supra, reported antiviral action of peptidomannan (KS-2) extracted from culture mycelia of Lentinus edodes. Both oral (PO) and intraperitoneal (IP) administration resulted in increased peak serum interferon titer which protected mice

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against viral infections. This was different from dextran phosphate (DP-40) (Suzuki et al., Proc. Soc. Exp. Biol. Med., 149: 1069 - 1075 (1975)) and 9-methylstreptimidone (9-MS) (Saito et al., Antimier. Agent & Chemotherapy, 10: 14 - 19 (1976)) which induced higher titers of interferon only if administrated intravenously (IV) or intraperitoneally (IP) to mice.

Anti-inflammatory activity of aloe vera gel has been widely reported by both oral testimonies and respected scientific journals. Rubel (Cosmetics and Toiletries, 98: 109 - 114 (1983)) discussed fully the possible mechanism of the anti-inflammatory effect of aloe gel. Ukai et al., (J. Pharmacology Dyn., 6: 983 - 990 (1983)) noted anti-inflammatory activity in polysaccharides extracted from fruit bodies of several fungi. The polysaccharides demonstrated significant inhibitory effect on carrageenan induced edema. One of the polymers, O-acetylated-D-mannan (T-2-HN), in addition demonstrated more marked inhibitory effect on scald hyperalgesia than phenylbutazone. Ukai et al., supra. The fact that the polysaccharide is said to be free from protein and lipids strongly suggests that the anti-inflammatory effect is due to the acetylated mannan only. Other researchers have also reported anti-inflammatory effects of complex polysaccharides (Saeki et al., Japan J. Pharmacology, 24: 109 - 118 (1974)), glycoproteins (Arita et al., J. Pharmacology, 24: 861 - 869 (1974)) and sulfated polysaccharides (Rocha et al., Biochemical Pharmacology, 18: 1285 - 1295 (1969)).

Literature reports of polysaccharides demonstrating pharmacological and physiological activities continue to flood the pages of well respected scientific journals. It is therefore, not illogical that the mucilaginous gel of the Aloe vera, which is essentially a polysaccharide, holds the secret to Aloe vera's medicinal properties. The discrepancies over whether the polysaccharide is a glucomannan, mannan, pectin, or some other composition are a result of chemical purification steps. By processing Aloe according to the present

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invention, a partially acetylated polymannose has been consistently isolated as the major polysaccharide having pharmacological activity. Yagi et al, (Planta Medica, 31: 17 - 20 (1977)) using a slightly modified extraction method, isolated acetylated mannan (aloe mannan) from Aloe arborescens Miller var. natalensis. Ovodova, (Khim. Prior. Soedin, 83: 93833 (1975)) however, earlier isolated pectin as the main component of the same aloe species.

SUMMARY OF THE INVENTION

The present invention is directed to a process whereby the active chemical substance in the aloe plant is physically extracted from whole aloe leaves. The chemical substance is substantially non-degradable and can be administered in a prescribed amount.

The present invention is also directed to the active chemical substance in the aloe plant in the form produced by the process described above. The active chemical substance has been found to be a substantially non-degradable lyophilized linear polymer of substantially acetylated mannose monomers. The mannose monomers are preferably bonded together by β (1-4) bonds. The active chemical substance has been measured, standardized and characterized by analytical chemistry techniques.

The term "active chemical substance" as used herein means the substance which is responsible for the wound healing and other beneficial properties of Aloe vera. The term "substantially non-degradable" as used herein means a product which decreases in molecular weight by less than 5 percent over a period of two years and a product which maintains more than 95 percent of its biological activity over a period of two years. The term "substantially acetylated mannose monomers" as used herein means partially or substantially completely acetylated mannose monomers.

The process of the present invention is one for extracting the active chemical substance in the aloe plant from a leaf of the aloe plant which process basically compri-

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ses at least the following steps:

(a) washing an aloe leaf in a bacteriacidal solution to remove substantially all surface dirt and bacteria;

(b) removing at least a first end portion from said washed leaf;

(c) draining, preserving and collecting anthraquinone rich sap from said cut and washed leaf;

(d) removing rind from said leaf to produce a substantially anthraquinone-free aloe gel fillet;

(e) grinding and homogenizing said substantially anthraquinone-free aloe gel fillet to produce substantially anthraquinone-free aloe juice having solubilized matter;

(f) adding a water soluble, lower aliphatic polar solvent to the aloe juice to precipitate the active chemical substance and thereby to form a heterogeneous solution;

(g) removing the water soluble, lower aliphatic polar solvent and the solubilized matter from the heterogeneous solution to isolate the precipitated active chemical substance; and

(h) drying the precipitated active chemical substance.

One skilled in the art will appreciate that one may obtain aloe juice having solubilized matter from aloe leaves in whatever manner possible, and then subject this juice to steps (f), (g) and (h) to extract the active chemical substance.

Indeed one skilled in the art will appreciate that instead of steps (b), (c) and (d), one may instead (b) crush the washed aloe leaves and (c) dialyze the crushed leaves chemically removing unwanted fractions, i.e., anthraquinones, minerals and acids and the rind to produce a substantially anthraquinone-free gel that may then be subjected to steps (e), (f), (g) and (h) to extract the active chemical substance.

One skilled in the art will also appreciate that instead of steps (b), (c), (d) and (e), one may instead crush

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the washed aloe leaves and extrude anthraquinone-rich aloe juice having solubilized matter and then subject the aloe juice to steps (f), (g) and (h) to extract the active chemical substance. The active chemical substance is effectively separated from anthraquinones and deleterious ions by this process since the anthraquinones and ions are water soluble and remain in the liquid solvent phase and do not precipitate.

One skilled in the art will also appreciate that instead of steps (b), (c), (d) and (e) one may instead grind the whole washed aloe leaves, filter out fibrous material, and homogenize the remainder to produce anthraquinone-rich aloe juice having solubilized matter. The aloe juice can then be subjected to steps (f), (g) and (h) to extract the active chemical substance. The active chemical substance is effectively separated from anthraquinones and deleterious ions by this process for the reasons noted above.

One skilled in the art will appreciate that an additional process for extracting the active chemical substance in the aloe plant from a leaf of the aloe plant comprises the following steps:

a) washing an aloe leaf in a bacteriacidal solution to remove substantially all surface dirt and bacteria;

b) removing rind from said leaf to produce an aloe gel fillet;

c) grinding and homogenizing said aloe gel fillet to produce aloe juice having solubilized matter;

d) adding a water soluble, lower aliphatic polar solvent to the aloe juice to precipitate the active chemical substance and thereby to form a heterogeneous solution;

e) removing the water soluble, lower aliphatic solvent and the solubilized matter from the heterogeneous solution to isolate the precipitated active chemical substance; and

f) drying the precipitated active chemical substance.

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As noted above, the active chemical substance is effectively separated from anthraquinones and deleterious ions by this process since the anthraquinones and ions are water soluble and remain in the liquid solvent phase and do not precipitate.

Removal of "substantially all surface dirt and bacteria" as used herein means (1) removal of dirt to the extent that remaining dirt is less than 0.1% by weight of the weight of the leaf and (2) killing such surface bacteria that the remaining surface bacteria are less than 100 count per gram of leaves.

Furthermore, my preferred process may further comprise the step of ultrafiltration in order to osmotically adjust the aloe juice or aloe vera fraction, or to reduce even further the levels of anthraquinones to less than 5 ppm, even down to less than 100 parts per billion by weight.

These steps enable the processor to use large or small leaves even less than one year old because the polymer size found in the mature leaves can be selected and processed out of smaller, immature leaves.

One of the advantages of the instant process is that damaged leaves previously considered unusable due to strong winds or poor collection techniques can be processed, and the undesirable contaminants can be dialyzed out.

The ultrafiltration (dialysis) step incorporates membrane technology that allows the selection of filters with different pore sizes, depending on the condition of the cut aloe leaves, that can accomplish any combination of the following:

- (1) A small pore size filter (preferably about 100 Daltons) that separates out water and salts from the Aloe vera gel, if needed.
- (2) Larger pore size filters (preferably about 500 Daltons) that can separate out the acids from the Aloe vera gel, if needed.
- (3) Even larger pore size filters (preferably about

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2000 Daltons) that can separate the yellow sap components from the Aloe vera gel, if needed.

- (4) And even larger pore size filters (preferably from about 10,000-100,000 Daltons that can size the gel matrix polymers, and divide them out by molecular weight.

A Romicon 4-column (Romicon Co., 100 Cummings Park, Woburn, MA 01801, Model No. HF4 SSS, Membrane Type PM50, Membrane Nos. H526.5 - 43 - pm50) as an ultrafiltration device is recommended.

As an additional preferred embodiment, the washing step of the process may comprise washing the substantially anthraquinone-free aloe gel fillet in a tumbler washer prior to grinding said fillet.

During all of the above-described processes of extracting the active substance from Aloe vera gel, minor amounts of organic and inorganic substances are found to co-precipitate with the product. A large fraction of the inorganic salts comprise calcium oxalate. The presence of inorganic salts such as calcium oxalate should be eliminated or at least minimized, for consistency of product and health reasons. It has now been surprisingly found that by adding an effective amount of a mineral acid to adjust the pH of the gel to about 3.0 to about 3.5 prior to addition of alcohol results in a product having a substantially lower level of oxalates and other inorganic salts. Accordingly, it is preferred in all of the above processes for extracting the active chemical substance in the aloe plant from a leaf of the aloe plant, that the pH of the gel be adjusted from about 3.0 to about 3.5 prior to the addition of alcohol.

Preferably in all the above processes for extracting the active chemical substance in the aloe plant from a leaf of the aloe plant, four volumes of the water soluble, lower aliphatic polar solvent are added to one volume of aloe juice to precipitate the active chemical substance. Preferred water soluble, lower aliphatic polar solvents are methanol, ethanol

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and propanol. The most preferred solvent is ethanol. One skilled in the art will recognize that other water soluble, lower aliphatic polar solvents may be substituted for the preferred solvents as long as the active chemical substance will precipitate therefrom.

It is preferred in the above extraction processes that the active chemical substance is allowed to precipitate from the water soluble, lower aliphatic polar solvent and aloe juice mixture for about four hours. It has been determined that allowing the mixture to precipitate for four hours gives the optimum yield of active chemical substance and that after four hours the precipitated active chemical substance begins to degrade. It has also been determined, however, that significant amounts of active chemical substance have been recovered after a precipitation period of 24 hours. One skilled in the art will appreciate that the optimum precipitation time period is dependent on ambient temperature and pressure as well as the nature of the water soluble, lower aliphatic polar solvent.

It is also preferred in the above extraction processes that the precipitated active chemical substance be dried by lyophilization rather than by oven drying since heat may aid in the hydrolysis or degradation of the active chemical substance.

In all of the above extraction processes, any fibrous material (cellulose) contained in the aloe juice is also precipitated by the water soluble, lower aliphatic polar solvent but is precipitated early with the addition of the solvent and is less dense than the active chemical substance. The fibrous material remains on the surface of the solvent after the active chemical substance has been allowed to settle and can therefore be removed quite easily. One skilled in the art will appreciate that one may instead filter the aloe juice to remove fibrous material prior to the addition of solvent.

More preferably, in all of the above processes, aloe leaves or whole plants may be collected from the field suf-

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ficiently clean to eliminate a washing step.

The dried, precipitated active ingredient, optionally, may be irradiated by gamma or microwave radiation whereby said active chemical substance is sterilized and preserved.

Accordingly, the present invention is believed to provide new and improved methods for the production of aloe vera products.

The present invention furthermore is further believed to provide improved methods for processing the leaf of the aloe vera plant in a manner which avoids the undesirable combination or mixture of distinctively characteristic portions of such plant leaf.

The present invention moreover is further believed to provide improved processes for preparing various extracts of the leaf of the aloe vera plant which minimizes the concentrations of undesirable components in the finished extracts.

The present invention also is believed to provide new and improved processes for preparing extracts of the leaf of the aloe vera plant which maximize the concentration of certain components characteristic of particular portions or segments of the leaf while minimizing or eliminating certain components characteristic of other portions or segments of the leaf.

The present invention also is believed to provide new and improved processes for preparing extracts from the leaf of the aloe vera plant which are low in concentration of yellow sap of the leaf.

The present invention finally is thought to provide a method for extracting the active chemical substance in Aloe vera gel. This chemical substance has utility as a non-toxic immune stimulating compound. The substance shall hereinafter be referred to as Carrisyn™ or acemannan. As mentioned above, acemannan has been found to be a substantially non-degradable lyophilized linear polymer of substantially acetylated mannose

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monomers which is standardized and characterized by analytical chemistry techniques.

DESCRIPTION OF THE FIGURES

Figures 1 and 2 show cut-away portions of an aloe vera leaf.

Figure 3 shows a schematic of a preferred leaf washing apparatus used in the process of the present invention.

Figure 4 shows a schematic of a preferred apparatus for sectioning and soaking of aloe vera leaves.

Figure 5 shows a schematic of preferred apparatus for the cutting of aloe sections into fillets and for coarse grinding.

Figure 6 shows a schematic of preferred apparatus for fine homogenization and filtering.

Figure 7 shows a schematic of a preferred use of dialysis equipment for fine separations of processed aloe material.

Figure 8 shows infrared spectra for two samples of Carrisyn™ under different pH conditions.

Figure 9 shows a differential thermogram of Carrisyn™.

Figure 10 shows a differential thermogram of Carrisyn™ contaminated with calcium oxalate.

Figure 11 shows a schematic for the characterization of Carrisyn™.

Figure 12 shows a size exclusion chromatogram of pullulan polysaccharide standards.

Figure 13 shows a size exclusion chromatogram of Carrisyn™.

Figure 14 shows an infrared spectrum of non-protease treated Carrisyn™.

Figure 15 shows an infrared spectrum of protease treated Carrisyn™.

Figure 16 shows a differential thermogram of

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Carrisyn™.

Figure 17 shows an HPLC chromatogram of a standard mixture of glucose, galactose, mannose and inositol.

Figure 18 shows an HPLC chromatogram of Carrisyn™.

Figure 19 shows a GLC chromatogram of a standard mixture of rhamnose, fucose, arabinose, xylose, mannose, galactose, glucose and inositol as their glycitol acetates.

Figure 20 shows a GLC chromatogram of Carrisyn™.

Figure 21 shows a standard curve of mannose/inositol ratio against amount of acemannan.

Figure 22 shows a total ion chromatogram of partially methylated and partially acetylated glycitol of Carrisyn™.

Figure 23 shows a mass spectrum of the partially acetylated glycitol of Carrisyn™.

Figure 24 shows a schematic for partially methylated mannitol acetates.

Figure 25 shows a schematic for fragment ions of methylated mannitol acetate under mass spectrometry analysis.

DETAILED DESCRIPTION OF THE PREFERRED EMBODIMENTS

It has been discovered and recognized that sub-portions of the yellow sap and internal gel matrix have characteristics which are unique to those sub-portions, the extracts therefrom therefore having potentially distinct uses from one another. A summary of these distinct portions, and some of their characteristics and potential uses is as follows:

<u>PORTION</u>	<u>SUB-PORTION</u>	<u>UTILITIES</u>
Yellow Sap	(1) Sediment	laxative, antifungal, antibiological, pesticidal and sunscreen
	(2) Supernatant	mucosal protective action, sunscreen
Internal Gel Matrix	(1) Mucilage	penetrant, hypo-allergenic, moisturizer

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(2) Gel Fillets	ulceroprotective, cell stimulative moisturizer, wound healing
(3) Interstitial Fibers	natural preservative, hemostat
(4) Residual Matrix	cell growth stimulant
Outer Rind	pesticidal insect re- pellent, paper pulp fiber

Based upon the aforementioned knowledge and recognition, and in order to optimize the quality and concentration of the desired components in the final extract depending upon the intended use thereof, the process of this invention is directed to the initial fractionating of the leaves of the aloe vera plant into the particular distinct portions and sub-portions defined above as well as the separation and isolation of particular components of such sub-portions defined above. The specific details and features of such processes will become more readily understood and appreciated from the following detailed description. The present invention is also directed to particular components that are isolated by the above-mentioned processes.

The products produced by the process of the present invention are preferably obtained from the outer, lower leaves of mature, outdoor grown aloe vera plants. A two year old plant is normally mature; however, the broader leaves of a four or five year old plant typically contain larger amounts of the desired extracts and are also easier to handle. Four or five year old leaves of Aloe barbadensis Miller grown in the Rio Grande Valley of Texas are most preferable. These leaves generally weigh about 1½ - 2½ pounds each. Depending upon the particular use or products that are desired, the leaves can be processed immediately after cutting them from the plant or they can be stored under appropriate conditions for varying time periods before they are processed.

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Additionally, concentration of various components of the leaf are affected by seasonal variations and the environmental conditions to which the leaves are subjected, all of which should be taken into consideration depending upon the specific intended use to which the plant extracts are to be directed.

The leaves should be pulled or cut from near the base of the plant, preferably without breaking or damaging any part of the leaf prior to processing. One preferably employs a small knife of less than 6 inches, e.g., a pocket knife and cuts the leaf at the base immediately above the stalk, and peels the leaf from the stalk, in order to prevent leakage of the clear cellular gel or contamination of the gel with the yellow sap. Any breakage or bruising of the leaf can result in the undesirable commingling of the distinct portions, and therefore component characteristics, of the leaf.

After removal from the plant, the leaves are normally cleaned by washing them with a mild scrubbing action or spraying with a suitable detergent solution (such as OLYMPIC POOL CHLOR 65™, distributed by York Chem Co., Dallas, Texas). In some instances, the cleaning takes place with the aid of soft brushes. After cleaning, the leaves are rinsed thoroughly in clean water to remove any vestige of any detergent solution.

The bottom white or light colored portion of each leaf and the tipmost portion thereof are removed by cutting carefully with a small sharp knife. These portions, essentially constituting the ends of the leaf, may be separately processed to obtain the yellow sap therefrom for those applications in which it is desired to produce products having components with the yellow sap characteristics referred to above.

The remaining portion of each aloe vera leaf is then crosscut into short segments, preferably one-half inch in length, and each segment is placed upright in an aqueous solution (preferably deionized water) which may be hypertonic, isotonic or hypotonic, resulting in the yellow sap draining

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from the segments. Alternatively, in those applications in which it is desired to collect the yellow sap for use in other preparations having the characteristics noted above, the segments may be placed upright in a dry collection container, preferably of stainless steel with a stainless steel wire mesh bottom to allow drainage and for water contact to allow the water to dialyze the leaves.

The segments are permitted to drain for approximately twenty to thirty minutes in this manner. The cut segments will eventually form a seal and cease draining. The yellow sap that is collected will then, upon standing for an appropriate period of time, separate into two sub-portions, namely sediment and supernatant, respectively. The yellow sap is useful for making a good sunscreen on intact skin (not broken skin) and provides the skin with an olive tan color, and is also useful for the manufacture of laxatives.

After completion of the procedures which remove the yellow sap from the cut leaf segments, the segments are then pared to form fillets utilizing any suitable equipment such as a wire (i.e., consumer cheese) slicer or paring blades (e.g., paring knife) to remove the outer rind or skin of the leaf segments and the layer that is immediately below such outer skin. The leaf segments may be frozen to facilitate this skinning procedure. After paring, what remains is the internal gel matrix portion (fillet), and this portion is inspected and hand cleaned to remove any adhering skin or discolored portions to eliminate any residual yellow sap therefrom. One uses a mild water spray, preferably deionized water (and free of alcohol) or submerges the gel matrix portion under flowing clean water, to facilitate removal of such yellow sap residue.

The resultant fillet (internal gel matrix) can then be drained for approximately an hour. During this draining procedure, a slimy coating usually forms on the surfaces of the gel matrix, this coating being collected by gravity or assisted by appropriate means such as centrifugation. This

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collected coating is the mucilage sub-portion referred to above.

The remainder of the gel matrix, in the form of gel matrix strips, may then be ground, shredded or blended to break up the interstitial fibers present therein, or the gel matrix strips may be forced through a wire mesh or filter screen in order to achieve liquefaction. This resultant substance may then be subsequently homogenized. Alternatively, the gel matrix strips may be frozen and thawed and subsequently mixed to produce a liquid substance with fibers (such substance constituting the gel fillets sub-portion referred to above). This substance can then be filtered to obtain the interstitial fibers sub-portion, leaving the residual matrix sub-portion referred to above.

The homogenized extract thus obtained typically has a pH of approximately about 4 to about 5, preferably about 4.

All steps in the process described are performed at about room temperature.

Figures 3-7 disclose in even further detail preferred embodiments of the instant process. Specifically, in Figure 3, there is disclosed apparatus for leaf washing. Aloe vera leaf washing equipment (Thompson Manufacturing Company, Harlington, Texas) A is utilized whereby leaves are first pre-soaked in vat 4. The whole leaves α are then placed by hand onto a conveyor belt 8 by which they are pulled underneath two brushes 9a and 9b. Conveyor belt 8 is driven via a chain 7 which is rotated on a second end by motor and pulley 6 which extends from a housing 5, also provided by Thompson Manufacturing Company. As the leaves are brushed and washed, upon passing through the second brush 9b the leaves are inspected at the end 10 of conveyor belt 8, by which the leaves are visually inspected and determined whether or not they are sufficiently clean. If the leaves are not sufficiently clean, they are placed into vat 12 for further washing; if the leaves are sufficiently clean, they are placed upon an upwardly moving conveyor B having steps 13 for which

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each individual leaf may be further washed with tap water by sprayers 11. The conveyor B is provided by Dallas Mill Wright Co. of Seagoville, Texas. The rinse sprayers 11 are provided by Key Plumbing, Seagoville, Texas. Stainless steel vat 12 is made of 316 stainless steel and is custom made by the National Sheet Metal Company of Dallas, Texas.

After washing, as indicated in Figure 4, the leaves are sectioned and soaked. After moving up through steps 13 of conveyor B, the clean, raw leaves α drop onto a tray 14 which is provided with a hole 15 for removal of trash. Tray 14 is part of a sectioning and soaking apparatus C of 316 stainless steel provided by the National Sheet Metal Company of Dallas, Texas. This equipment is custom fabricated. On tray 14 the leaves are manually sectioned at both ends with the tips and butts disposed through the hole 15 into a trash receptacle (not shown). The cut leaves β are then stacked into any one of a number of baskets 16 of stainless steel each having a wire mesh bottom made of stainless steel. Then these baskets 16 are placed in a stainless steel track which forms the top part of a trapezoidal funnel 17 by which yellow sap is drained from the bottom of the leaves through the baskets, falling into the bottom portion of the funnel 17. The yellow sap is periodically removed and is kept frozen for storage. The yellow sap draining step takes about 30 minutes.

After this step, the cut leaves β, still retained in basket 16 are manually moved to water bath 18 at positions closest to trapezoidal funnel 17. In countercurrent flow, water comes into bath 18 through inlet water pipe 19, at a point farthest removed from trapezoidal funnel 17 and is subsequently removed through exit water pipe 20 at a position closest to trapezoidal funnel 17. Trays are gradually moved manually through the water bath in a direction away from trapezoidal funnel 17, and the washing step whereby the baskets remain in water bath 18 takes approximately one hour.

After washing, the baskets are placed on tray 21 for drying, which lasts for only a few minutes. Again, the entire

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assembly in Figure 4, pertaining to the baskets, including wire mesh, yellow sap draining and auto washing equipment is made of 316 stainless steel custom fabricated by the National Sheet Metal Company, Dallas, Texas.

After washing, the cut leaves Ø stacked in basket 16 on tray 21 are then moved to an area for further sectioning into fillets, as shown in Figure 5. Rind 22 is removed from the fillets so that only substantially clear material remains. The rest of the rind 22 is discarded. Fillets Ø are then placed into trough 23 which feeds to a rough coarse grinder D. The trough 23 is manufactured of 316 stainless steel by the National Sheet Metal Company of Dallas, Texas. The grinder is a Model No. 5150-N-Sink-erator (Watson Food Service Industries, Inc., Dallas, Texas). After rough coarse grinding through grinder D, the processed material Ø' emerging from the grinder passes through to a portable tank E of approximately 100 gallon capacity which comprises a vertical single shell tank of 316 stainless steel (Perma-San, Manufacturer, distributed through Van Tone, Inc., Dallas, Texas). Coarse ground fillet Ø' is agitated in tank E by a Perma-San agitator (Model No. AAPH2) also distributed through Van Tone, Inc. of Dallas, Texas.

After rough homogenization, material from tank E is taken to a separate staged area for fine homogenization and (optional) filtering. In Figure 6, material from tank E is pumped through pump 26 (centrifuge pump provided by Crepaco, Inc., Dallas, Texas, Model No. 4V-81, of stainless steel) driven by Reliance Motor Model No. B76Q2139M-VF of the Reliance Electric Company, Cleveland, Ohio into a fine homogenizer F (Crepaco, Inc., Dallas, Texas, Model No. 3DD13). After fine homogenization the material is transported to a large 1,000 gallon vertical single shell mixing tank G made of 316 stainless steel (Perma-San, Manufacturer, distributed to Van Tone, Inc. of Dallas, Texas). The finely ground fillet △ is agitated by a Perma-San agitator 26a (Perma-San, Model No. AAPH2, distributed through Van Tone, Inc., Dallas, Texas).

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Material A in tank G may be removed and sent to pump 27a which forms one unit with filter 27b for removal of pulp through discard line 28. Pump 27a and filter 27b form the part of a Lomart diatomite filter (Model No. 99-2138 distributed by Allen Recreation Company, Dallas, Texas). Filtered material is then pumped into tank H, which like tank E, can be provided with a lid 25.

In Figure 7, finely ground fillet A, is partially filtered, is stirred by mixer 29 and pumped through pump 30 into a dialyzer. Mixer 29 is a Perma-San agitator (Model No. AAPH2) distributed through Van Tone, Inc. Company, Dallas, Texas. Pump 30 is a Superior stainless steel Model SCS45 process pump (Superior Stainless Company, Delavan, Wisconsin). Attached to process pump 30 is motor 30a of 3 horsepower made by Baldor Motor of 4450 rpm (Cat. No. CM3559T of the Baldor Electric Company, Ft. Smith Arkansas). Material pumped through pump 30 passes through dialysis unit J, (a Romicon Model HF4SSS ultrafiltration system made by Romicon Inc., Woburn, Massachusetts) having 4 filters 31 (not shown) each filter housed in filter housing 32. Material passes vertically to a point where portions can be removed through separation lines 34 and into separation discharge lines 35. Other material not separated is recycled through recycle return line 33 back into the dialysis unit, or in the alternative, through separation return line 36 back into the vat I.

Depending on the fraction of aloe desired and end product sought, the desired material can either be obtained through separation discharge line 35 after processing or in vat I. For example, if excess water and minerals need to be removed, a small pore size ultrafilter can be used to separate out the water and minerals which are discharged through line 35 and the desired aloe fraction is returned to vat I. This process can be repeated until the desired amount of salt and water is removed from the product contained in vat I simply by circulating it through the dialysis unit. This process can include more than one dialysis step. For example, as pre-

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viously described, the salts, low molecular weight acids and unwanted anthraquinones can be removed in a first dialysis step. The unwanted material is discharged through separation discharge line 35 and the desired fraction is returned to vat I. This step is accomplished by using ultrafilters obtained from Romicon with 10,000 Dalton pores. Next, the 10,000 Dalton ultrafilters are replaced with 50,000 Dalton ultrafilters also obtained from Romicon and the dialysis process is repeated. The dialysis process now separates the gel matrix polymers into two fractions; a first fraction consists of gel matrix polymers ranging in size from 10,000 to 50,000 Daltons and is discharged through separation discharge line 35, a second fraction consists of gel matrix polymers greater than 50,000 Daltons in size and is returned to vat I. This process can last from minutes to hours depending upon the amount of salt and water that is needed to be removed from a given product.

In Figure 7, separation return line 36 is made of Tygon tubing (food grade) provided by the Texas Rubber Supply Company, Dallas, Texas. The separation discharge line 35 is a 316 stainless steel pipe distributed by Van Tone, Inc., Dallas, Texas.

The residual matrix sub-portion may then be treated to separate, isolate and purify the active chemical substance, acemannan, in Aloe vera gel. To separate acemannan from the residual matrix sub-portion, an excess of a water soluble, lower aliphatic polar solvent is added to the residual matrix sub-portion. Acemannan then begins to precipitate from this mixture. The solution is allowed to settle for a sufficient period of time to allow as much active ingredient to precipitate out of solution as possible but not so long that the acemannan begins to degrade. After this period of time the supernatant is decanted or siphoned off without disturbing the settled precipitate. The precipitate and remaining solution are then placed in an appropriate centrifuge apparatus and the precipitate is collected into a pellet. After centrifugation,

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the supernatant is decanted and discarded. Optionally, the pellet is washed with a fresh aliquot of the water soluble, lower-aliphatic polar solvent and collected again and the supernatant is again discarded. The pellet is then lyophilized to dryness and allowed to dry overnight. The resulting product is a substantially non-degradable lyophilized form of acemannan. The resulting product may be ground to form a powder. Preferably, a suitable acid is added to Aloe vera gel prior to the addition of the lower aliphatic polar solvent. The acid is added to solubilize calcium oxalate contained in the Aloe vera gel in order to facilitate its removal. The acid is preferably added at a strength sufficient to solubilize the calcium oxalate impurity while not degrading the acemannan polymer chain.

An alternate and preferred process for separating and isolating acemannan includes the following steps.

Leaves from mature, outdoor grown aloe vera plants are pulled or cut from near the base of the plant, preferably without breaking or damaging any part of the leaf prior to processing. The leaf is preferably cut at the base of the plant immediately above the stalk and is peeled from the stalk, in order to prevent leakage of the clear cellular gel or contamination of the gel with the yellow sap.

After removal from the plant, the butt and tip portions of the leaves are removed and the cut leaves are pared to form fillets as described above in the fractionation process.

The resultant fillet (internal gel matrix) is then ground, shredded or blended to break up the interstitial fibers present therein or the internal gel matrix may be forced through a wire mesh or filter screen in order to achieve liquefaction. The resultant liquefied internal gel matrix is then homogenized. The homogenized extract thus obtained typically has a pH of approximately about 4 to about 5, preferably about 4. The homogenized extract is then filtered to remove the interstitial fibers. The homogenized

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and filtered extract may then be treated in the identical manner as the residual matrix sub-portion, referred to immediately above, to separate and isolate acemannan.

An additional alternate and preferred process for separating and isolating acemannan includes the following steps.

Leaves from mature, outdoor grown aloe vera plants are pulled or cut from near the base of the plant, preferably without breaking or damaging any part of the leaf prior to processing. The leaf is preferably cut at the base of the plant immediately above the stalk and is peeled from the stalk, in order to prevent leakage of the clear cellular gel or contamination of the gel with the yellow sap.

The leaves may then be crushed by suitable means, for example a "Thompson Aloe Extruder" made by Thompson Manufacturing Company, Harlingen, Texas, to extrude aloe juice. The extruded aloe juice may then be treated in the identical manner as the residual matrix sub-portion, referred to above, to separate and isolate acemannan.

All steps in the processes described are performed at about room temperature except for the lyophilization step which is preferably performed at about -50°C.

Various modifications of the disclosed processes and compositions of the invention, as well as alternative modifications, variations and equivalents will become apparent to persons skilled in the art upon reading the above general description. The following examples are illustrative only and are not intended to limit the scope of the appended claims, which cover any such modifications, equivalents or variations.

Example 1

Process for Separating and Isolating acemannan.

A. PRELIMINARY WORK:

1. Previously cleaned tanks, mixers and fittings were sanitized with 50% isopropyl alcohol (IPA) solution and

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rinsed free of IPA with hot deionized water.

2. Pumps and attached hoses were drained with 5% "HTH" chlorine swimming pool solutions, then flushed with water.

3. The pumps and attached hoses were sanitized with 50% isopropyl alcohol solution. Pumps and attached hoses were flushed with hot deionized water until free of isopropyl alcohol.

4. A homogenizer and attached hoses and pumps were sanitized with 50% isopropyl alcohol solution. The homogenizer and attached hoses were flushed with hot deionized water until free of isopropyl alcohol.

Aloe barbadensis Miller leaves collected from the Rio Grande Valley were transferred in a refrigerated truck at 40 to 45°F within eight hours after harvest and stored under refrigeration at 40 to 45°F until processed to reduce degradation.

Twenty to sixty pounds of stored leaves were then placed in a prewash bath of an aqueous solution of calcium hypochlorite at room temperature substantially to remove surface dirt from the leaves and kill surface bacteria on the leaves. The aqueous solution of calcium hypochlorite was prepared by adding approximately 0.125 grams of 98% calcium hypochlorite to one liter of water to produce a solution containing 50 ppm of free chlorine. The leaves remained in the prewash bath for a period of approximately five minutes..

Next, the substantially dirt and bacteria free leaves were placed on the horizontal conveyor belt of a "Thompson Aloe Washer" made by Thompson Manufacturing Company, Harlingen, Texas. The "Thompson Aloe Washer" washed the leaves with room temperature water to remove the surface dirt and aqueous solution of calcium hypochlorite from the leaves. Again, the leaves were visually inspected and hand scrubbed as necessary to remove any surface dirt remaining on the leaves. Such leaves were then rinsed with room temperature water.

The tip and butt portion was then removed from each

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leaf and the leaves were placed in stainless steel basket-type containers, placed together on top of a funnel shaped stainless steel collector, each container having a mesh bottom. Yellow sap was allowed to drain from the leaves for approximately 30 minutes. The yellow sap passed through the mesh bottom of the stainless steel basket and was collected in a funnel shaped collector.

The stainless steel basket type containers containing the aloe leaves were removed from the collector, and were then submerged in a second stainless steel vessel comprising a room temperature water bath of continuously horizontally flowing rinse water moving counter-current to the containers which are slowly moved by hand from one end of the vessel to the other, for approximately thirty minutes to one hour. This allows the yellow sap to drain further from the leaves. The leaves were allowed to soak in this solution for 30 minutes.

The leaves were then removed from this solution and the rind was removed from each leaf with a sharp knife or wire cheese slicer to produce an aloe gel fillet from each leaf portion. The aloe gel fillets were visually inspected and any contaminated aloe gel fillets or fillet part detected by a characteristic yellowish discoloration were discarded. The total mass of uncontaminated aloe gel fillets was 20 to 60 percent the starting leaf mass depending on the leaf size and condition.

The uncontaminated aloe gel fillets were then placed in a 750 seat restaurant set stainless steel garbage disposal unit which coarse ground the fillets to an average particle size of the consistency of a thick, but freely flowing (coarse homogenized) liquid. The stainless steel garbage disposal unit was made by the N-sink-erator Division of Emerson Electric Co., Racine, WI, Model No. SS-150-13, Serial No. 115132.

The coarse ground aloe gel fillets then passed to a 100 gallon stainless steel holding vat. The holding vat was

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made by Process Equipment Corp. of Belding, Michigan, Model No. 100 gallon OVC, Serial No. 40865-3.

From the holding vat, the coarse ground aloe gel fillet solution was pumped to a homogenizer. The homogenizer was made by Crepaco Food Equipment and Refrigeration, Inc., of Chicago, Illinois, Serial No. 04-03. The homogenizer was of a type typically used in dairy processes for the homogenization of milk. The coarse ground aloe gel fillet solution was finely homogenized under a pressure of about 1,500 psi.

From the homogenizer the finely homogenized aloe gel fillet solution was pumped to a stainless steel storage tank. The storage tank was made by Process Equipment Corp. of Belding, Michigan, Model No. 1000 gallon OVC, Serial No. 40866-2. The total mass of the homogenized aloe gel fillet solution was 20 to 60 percent of the starting leaf mass. Then, if necessary, the homogenized product was dialyzed using ultrafiltration.

The finely homogenized aloe gel fillet solution was then filtered to remove the interstitial fibers using a Leslie's Diatomaceous Earth Filter Model DE-48. The interstitial fibers themselves instead of diatomaceous earth were used as the filter media, the fibers being supported by a nylon mesh cloth filter support. The gel was pumped through the filter for several minutes before opening the exit port so that a sufficient amount of fibers could build up to serve as the filter media.

Twenty gallons of the filtered aloe gel fillet solution was then pumped into a 100 gallon tank and 80 gallons of 190 proof undenatured ethanol (Ethyl Alcohol, 190 proof, U.S.P., punctilious, 54 gal. Batch I.D. #CT185JO4 available through U.S. Industrial Chemicals, Co., P. O. Box 218, Tuscola, Illinois 61953) was added to the aloe gel fillet solution. The solution was stirred for 20 to 30 minutes using a Perma-San propeller agitator model #AAPGF-4A, Process Equipment Corp., Belding, Michigan.

The alcohol-gel solutions were then immediately

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transferred to several 11 quart pans 10½" diameter and 8" high of 18-8 stainless steel (Bloomfield Industries Inc., Chicago, Illinois, available through Watson Food Service Equipment and Supplies, 3712 Hagger Way, Dallas, Texas).

The alcohol-gel solutions were then allowed to settle for approximately four hours.

The clear liquid supernatant was then decanted or siphoned off using care not to disturb the precipitate that had settled on the bottom of the pans. The precipitate and remaining solutions were then transferred to four 1 pint stainless steel centrifuge buckets with about 500 g. of precipitate and remaining solution being transferred to each bucket. The buckets were then spun at 2000 x g for about 10 minutes using an IEC Centra-7 Centrifuge (International Equipment Co., 300 2nd Avenue, Needham Heights, Massachusetts 02194 available through: American Scientific Products, P.O. Box 1048, Grand Prairie, Texas 75050).

After centrifugation the supernatants were decanted and discarded. The pellets were then washed with fresh 190 proof, undenatured ethanol and collected again at 2000 x g for about 10 minutes. After spinning again the supernatant was discarded.

The pellet was then transferred to several 600 ml. VIRTIS lyophilization jars and swirled in liquid nitrogen until frozen. The lyophilization jars were then attached to a lyophilization apparatus consisting of a Welch Duo-seal Vacuum Pump (Model No. 1402, available from Sargeant-Welch, P.O. Box 35445, Dallas, Texas 75235), a Virtis Immersion Coil Cooler (Model No. 6205-4350 Cooler in acetone bath) and a Virtis 18 Port Vacuum Drum Manifold (Model No. 6211-0350). All Virtis equipment is available from American Scientific Products, P.O. Box 1048, Grand Prairie, Texas 75050. The lyophilization drum was filled with acetone that was maintained at -50°C.

The samples were lyophilized to dryness overnight and were then weighed on a Mettler AE 163 balance. The samples remaining consisted of substantially non-degradable

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lyophilized Carrisyn™. The yield from 20 gallons of aloe vera gel was approximately 145-155 g. of Carrisyn™.

EXAMPLE 2

Process for Separating and Isolating acemannan.

Aloe barbadensis Miller leaves collected from the Rio Grande Valley were transferred in a refrigerated truck at 40 to 45°F within eight hours after harvest and stored under refrigeration at 40 to 45°F until processed to reduce degradation.

The tip and butt portion was removed from each leaf. The rind was then removed from each leaf with a sharp knife or wire cheese slicer to produce an aloe gel fillet from each leaf portion.

The aloe gel fillets were then placed in a 750 seat restaurant set stainless steel garbage disposal unit which coarse ground the fillets to an average particle size of the consistency of a thick, but freely flowing (coarse homogenized) liquid. The stainless steel garbage disposal unit was made by the N-sink-erator Division of Emerson Electric Co., Racine, WI, Model No. SS-150-13, Serial No. 115132.

The coarse ground aloe gel fillets then passed to a 100 gallon stainless steel holding vat. The holding vat was made by Process Equipment Corp. of Belding, Michigan, Model No. 100 gallon OVC, Serial No. 40865-3.

From the holding vat, the coarse ground aloe gel fillet solution was pumped to a homogenizer. The homogenizer was made by Crepaco Food Equipment and Refrigeration, Inc., of Chicago, Illinois, Serial No. 04-03. The homogenizer was of a type typically used in dairy processes for the homogenization of milk. The coarse ground aloe gel fillet solution was finely homogenized under a pressure of 1,500 psi.

From the homogenizer the finely homogenized aloe gel fillet solution was pumped to a stainless steel storage tank. The storage tank was made by Process Equipment Corp. of Belding, Michigan, Model No. 1000 gallon OVC, Serial No.

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40866-2. The total mass of the homogenized aloe gel fillet solution was 20 to 60 percent of the starting leaf mass. Then, if necessary, the homogenized product was dialyzed using ultrafiltration.

The homogenized gel was then filtered to remove the interstitial fibers using a Leslie's Diatomaceous Earth Filter Model DE-48. The interstitial fibers themselves instead of diatomaceous earth were used as the filter media, the fibers being supported by a nylon mesh cloth filter support. The gel was then pumped through the filter for several minutes before opening the exit port so that a sufficient amount of fibers could build up to serve as the filter media.

Twenty gallons of the filtered gel was then pumped into a 100 gallon tank and 80 gallons of 190 proof undenatured ethanol (Ethyl Alcohol, 190 proof, U.S.P., punctilious, 54 gal. Batch I.D. #CT185J04 available through U.S. Industrial Chemicals, Co., P.O. Box 218, Tuscola, Illinois 61953) was added to the aloe gel fillet solution. The solution was stirred for 20 to 30 minutes using a Perma-San propeller agitator model # AAPGF-4A, Process Equipment Corp., Belding, Michigan.

The alcohol-gel solutions were then immediately transferred to several 11 quart pans 10½" diameter and 8" high of 18-8 stainless steel (Bloomfield Industries Inc., Chicago, Illinois, available through Watson Food Service Equipment and Supplies, 3712 Hagger Way, Dallas, Texas).

The alcohol-gel solutions were then allowed to settle for approximately four hours.

The clear liquid supernatant was then decanted or siphoned off using care not to disturb the precipitate that had settled on the bottom of the pans. The precipitate and remaining solutions were then transferred to four 1 pint stainless steel centrifuge buckets with about 500 g. of precipitate and remaining solution being transferred to each bucket. The buckets were then spun at 2000 x g for about 10 minutes using an IEC Centra-7 Centrifuge (International

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Equipment Co., 300 2nd Avenue, Needham Heights, Massachusetts 02194 available through: American Scientific Products, P. O. Box 1048, Grand Prairie, Texas 75050).

After centrifugation the supernatants were decanted and discarded. The pellets were then washed with fresh 190 proof, undenatured ethanol and spun down again at 2000 x g for about 10 minutes. After spinning again the supernatant was discarded.

The pellet was then transferred to several 600 ml. VIRTIS lyophilization jars and swirled in liquid nitrogen until frozen. The lyophilization jars were then attached to a lyophilization apparatus consisting of a Welch Duo-seal Vacuum Pump (Model No. 1402, available from Sargeant-Welch, P.O. Box 35445, Dallas, Texas 75235), a Virtis Immersion Coil Cooler (Model No. 6205-4350 Cooler in acetone bath) and a Virtis 18 Port Vacuum Drum Manifold (Model No. 6211-0350). All Virtis equipment is available from American Scientific Products, P.O. Box 1048, Grand Prairie, Texas 75050. The lyophilization drum was filled with acetone that was maintained at -50°C.

The samples were allowed to dry overnight and were then weighed on a Mettler AE 163 balance. The samples remaining consisted of substantially non-degradable lyophilized Carrisyn™. The yield from 20 gallons of aloe vera gel was approximately 145-155 g. of Carrisyn™.

EXAMPLE 3

Standard Laboratory Scale Process for Separating and Isolating Carrisyn™

Approximately 50 pounds of Aloe barbadensis Miller leaves were washed with water and brushed to remove dirt, dried latex and other contaminants. The outer cuticle of each leaf was then removed and the whole fillets were placed in a large beaker (on ice).

In 1.5 liter batches a Waring blender was loaded with the whole Aloe fillets. The fillets were blended at high

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speed twice for two minutes at room temperature. The blended fillets were then cooled at 4°C to allow foam generated during blending to settle.

The blended Aloe juice was then filtered through four layers of cotton (Cleveland Cotton) to remove any fibrous cellulosic pulp. The filtrate was then passed through six layers of cotton and approximately 4 liters of Aloe juice was collected.

The Aloe juice was then placed in a large five gallon stainless steel container. To the filtered juice was added 16 liters of chilled ethanol (Fisher Ethanol reagent grade, Cat. No. A995). The ethanol was added slowly while stirring the Aloe juice. A flocculent precipitate formed and the mixture was stirred for 15 to 30 minutes and was allowed to settle at room temperature for about two hours.

The supernatant was then decanted off and the pellet was placed in a small blender to which one liter of deionized water was added. This mixture was blended for several minutes at low speed to wash the pellet and was then placed in an eight liter nalgene container. To this mixture was added four more liters of ethanol and the mixture was stirred for 30 minutes. The precipitant that formed was allowed to settle for about two hours.

The majority of the supernatant was then decanted off and the resultant pellet was centrifuged at 2000 x g for 20 minutes at room temperature to pellet the precipitant for easy decanting of the remaining solvent.

The pellet was then placed in a lyophilization flask and lyophilized overnight in a Virtis lyophilizer.

The lyophilized powder weighed 10.9 g. The percent yield was 0.273% or 2.73×10^{-3} g/ml.

EXAMPLE 4

The Reduction Of Calcium Oxalate Contaminant
In Aloe Vera Extract - Acemannan

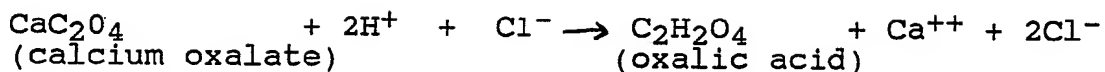
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During the process of extraction of acemannan from Aloe vera gel by alcohol, minor amounts of organic and inorganic substance are found to co-precipitate with the product. A large fraction of the inorganic salts and indeed the major contaminant in the alcohol extraction of Aloe vera gel to acemannan is calcium oxalate. The presence of calcium oxalate is confirmed by optical microscopy, infrared spectroscopy and thermogravimetric analysis. Although the quantity of calcium oxalate may vary from batch to batch, calcium oxalates accounting for about 30% of the total extract by weight have been observed. The purpose of the process described in Example 4 is to reduce the oxalate content in acemannan. Calcium oxalate is very insoluble in water and alcohol. By treating aloe gel with alcohol, the oxalate in the Carrisyn precipitate is concentrated. This concentration of oxalate is demonstrated by the strong infrared absorption of Carrisyn between $1600-1587\text{ cm}^{-1}$ due to the carbonyl asymmetric stretch of the oxalate and the high ash content of Carrisyn from thermogravimetric analysis. Since the acemannan is the active substance in Carrisyn, inorganic salts such as calcium oxalate should be eliminated or at least minimized for consistency of product quality and for health reasons.

One procedure for separating the oxalates and other inorganic salts would be by membrane dialysis. However, this method has many drawbacks and disadvantages. The process is very time consuming since the crude Carrisyn must be rehydrated, re-extracted in alcohol and freeze dried again. Most importantly, unless the product is preserved during the dialysis step, microbial degradation and spoilage will result in very inactive Carrisyn.

Generally, carboxylic acid salts are converted to their corresponding acids when treated with dilute mineral acids. By treating calcium oxalate with hydrochloric acid, oxalic acid and calcium salts of oxalic acid would be produced according to the following mechanism:

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Oxalic acid is highly soluble in water and alcohol and is therefore preferentially extracted into the water/alcohol mixture. The result is that the Carrisyn has a much lower level of oxalates and other inorganic salts.

About 2 liters of Aloe vera gel previously homogenized at 500 psig and filtered through a new swimming pool filter was collected for this Example. An initial pH of the gel was taken after stirring. Gel samples of known weight were placed into several 600 mL beakers and the pH adjusted as appropriate with a suitable acid, here a dilute mineral acid (preferably 6N hydrochloric acid) or concentrated sodium hydroxide solution, as follows:

Sample #	0	1	2	3	4	5	6	7	8
Weight(gel)	100.	100.	99.3	100.	100.	99.7	100.	99.9	99.5
pH	4.56	4.07	2.45	3.35	5.09	10.6	6.16	7.0	8.58

After the pH of each sample was adjusted four (4) volumes of SDA-3A ethanol were added as soon as possible to minimize the time the gel would be under the stipulated pH condition. After stirring the mixture for less than 2 minutes, each mixture was allowed to stand for about 3-4 hours. Each extract was then centrifuged, freeze dried and weighed to determine the yield under the various pH conditions.

The infrared spectra of the nine solid Carrisyn samples were obtained from a disc made by mixing an appropriate amount of the substance in potassium bromide. Each disc was scanned from 4000 to 400 cm^{-1} (wave numbers) on an IBM FT-IR spectrometer. The spectra of the Carrisyn samples at different pH conditions were compared qualitatively.

Thermogravimetric Analysis:

The thermal weight loss of Carrisyn has a definite finger print profile. About 10 milligrams of each sample was

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heated in a Mettler thermoanalyzer from 25°C to 780°C at a rate of 20°C/min. Nitrogen gas atmosphere was used until a temperature of 600°C was reached, followed by oxidizing environment to 780°C and then each sample was allowed to remain at this temperature for 2 minutes. From the weight loss thermogram, the moisture content, the carbohydrate, oxidizable carbon skeleton, oxalate and ash contents were determined.

Results and Discussion:

The most important factor with an acid treatment of Carrisyn is whether the physiochemical properties of the substance have been adversely affected by the process. A suitable acid must be selected: (i) an acid capable of achieving the proper pH range (from about 3.0 to about 3.5) in reasonable volumes, (ii) will not react adversely with beneficial components (polydisperse acemannan), the solvent mixture (ethanol) and the holding vessels and equipment; and (iii) additionally, chosen in such a concentration as not to degrade the acemannan chain. Many dilute mineral acids and organic acids in higher concentrations are suitable, although a non-oxygenated mineral acid is most suitable as esterification and degradation are minimized. With the selection of a suitable acid, e.g., 6N hydrochloric acid, rather than having an adverse affect, the quality of Carrisyn appears to improve appreciably at lower pH conditions. For example, the same concentration (w/v) of Carrisyn gave a more viscous, rehydrated solution when acid treated. A viscous solution signifies good product. Size exclusion chromatographic separation of the solution demonstrated the same chromatographic profile as an untreated Carrisyn. Therefore, under the conditions applied, degradation was not observed.

The following results demonstrate that the yield of Carrisyn as represented by total solid in a unit volume of gel is reduced after acid treatment:

Sample #	0	1	2	3	4	5	6	7	8
----------	---	---	---	---	---	---	---	---	---

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pH	4.56	4.07	2.45	3.35	5.09	10.56	6.16	7.0	8.68
Yield%	0.26	0.22	0.12	0.14	0.27	0.23	0.25	0.26	0.24

However, subsequent analyses of the products by infrared spectroscopy and thermogravimetric analysis reveal that the high yield clearly correlated to oxalate and ash content.

Infrared spectroscopy may be used extensively to characterize Carrisyn both qualitatively and quantitatively. The increase in the absorption peak of the acetyl group located about 1740 cm^{-1} and the decrease of the oxalate peak about 1590 cm^{-1} of Carrisyn at lower pH demonstrate a reduction of the oxalate content (Figure 8).

Thermogravimetric analysis gives a characteristic weight loss profile with temperature. Generally, pure Carrisyn gives 3 major peaks on the differential thermogram namely; (a) moisture $30\text{--}100^\circ\text{C}$, (b) carbohydrate (acemannan) $200\text{--}400^\circ\text{C}$, and (c) oxidizable carbon skeleton $600\text{--}630^\circ\text{C}$ (Figure 9). On the other hand, Carrisyn contaminated with calcium oxalate for example exhibits 2 more peaks located between $456\text{--}500^\circ\text{C}$ and $650\text{--}720^\circ\text{C}$ with high ash content (Figure 10).

Tables 1A - E below present thermogravimetric analyses of Carrisyn at various pH conditions.

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TABLE 1A
THERMOGRAVIMETRIC ANALYSIS
OF CARRISYN pH 4.56

SAMPLE ID	WT (mg)	%H ₂ O	% RESIDUE	% CARBO-HYDRATE	OXALATE	
					425 540 C	631 750 C
70302	9.7220	6.9122	17.3940	49.9485	10.3680	12.14
70307	8.8020	6.6803	18.0750	51.1364	9.7364	11.64
70311	10.1150	6.7721	18.9030	51.3980	8.6604	11.36
70312	9.1790	6.6347	18.4330	52.4889	8.7700	10.89
70313	8.9810	5.2667	18.5390	53.0124	8.9522	10.85
70314	9.4530	4.7287	18.5130	52.0994	9.4785	12.19
70316	9.6000	5.4271	17.7810	52.5517	9.1771	12.19
70320	9.9120	5.6094	17.0600	54.7322	8.6259	10.62
AVERAGE (X)	9.4705	6.0039	18.0873	52.1709	9.2211	11.49
STD DEV	0.4570	0.8391	0.6325	1.4205	0.6087	0.65
N = 8						

TABLE 1B
THERMOGRAVIMETRIC ANALYSIS
OF CARRISYN pH 3.60

SAMPLE ID	WT (mg)	%H ₂ O	% RESIDUE	% CARBO-HYDRATE	OXALATE	
					425 540 C	631 750 C
70315	9.2100	6.9164	7.7959	72.4104	5.6243	4.082
70321	10.7180	6.5777	9.0035	70.1716	6.4751	4.879
70323	9.3070	6.0277	7.9295	73.5680	5.1467	3.900
70324	9.0919	5.2690	6.9739	74.6780	5.0820	3.707
70326	9.6950	7.0397	10.1080	70.0355	6.4879	5.796
70327	9.1490	5.7493	9.5639	69.4942	6.6455	5.224
70328	9.4680	5.5450	8.3333	68.8317	9.5902	4.393
70329	9.3140	6.4204	8.6859	71.6130	5.5937	4.402
AVERAGE (X)	9.4940	6.1939	8.5491	71.3503	6.3307	4.548
STD DEV	0.5303	0.6478	1.0118	2.0730	1.4527	0.708
N = 8						

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TABLE 1C
THERMOGRAVIMETRIC ANALYSIS
OF CARRISYN pH 3.20

SAMPLE ID	WT (mg)	%H ₂ O	% RESIDUE	% CARBO- HYDRATE	OXALATE	
					425 540 C	631 750 C
70330	9.3260	3.5814	6.2621	77.1819	5.7152	3.506
70331	9.8330	4.6273	4.5764	80.7280	4.9425	1.871
70334	10.8360	6.1831	5.0388	77.6395	4.8819	2.685
70401	8.1040	7.8196	4.8618	80.3070	4.9975	2.023
70402	9.0650	4.0927	4.6222	79.4380	6.8285	1.522
70403	9.9490	4.2014	4.3220	80.1990	4.7945	2.472
70405	8.8930	6.0497	4.8353	79.1636	4.4305	2.069
70406	9.8350	4.3213	4.0874	80.9156	4.7788	1.9420
70419	9.0080	5.3619	2.1758	81.8822	4.7069	1.7651
70423	9.8510	6.6998	3.9285	79.8803	4.7102	2.0506
70426	9.7880	5.5885	4.1479	78.8307	3.9947	1.2873
70427	10.1010	6.7815	3.2670	79.5866	4.5144	2.0190
70428	10.7790	3.6181	2.6069	80.1190	4.4717	2.7368
70431	9.9030	5.9881	4.1099	78.4407	4.5542	2.3831
70433	8.3980	6.2277	3.5127	79.5073	4.5844	2.2982
70434	9.7360	7.1385	4.6188	77.1878	4.5604	2.5781
70440	10.0850	5.3148	3.1730	81.4331	4.4522	1.9534
70442	9.9470	6.2933	3.3779	81.0805	4.4234	1.2567
70504	9.9740	6.2933	4.7925	77.1410	4.7323	2.8173
70505	9.1440	5.9165	3.6308	80.5337	4.4182	1.4545
70506	10.2590	5.5659	3.4214	78.8573	5.9460	1.9300
70514	10.2060	7.7406	5.5164	74.4461	5.2714	4.0956
AVERAGE (X)	9.6827	5.7172	3.8581	79.2956	4.8504	2.2145
STD DEV	0.6851	1.2332	1.3386	1.7546	0.6198	0.6831
N = 22						

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TABLE 1D
THERMOGRAVIMETRIC ANALYSIS
OF RECYCLED CARRISYN
20 MINS AGITATION

SAMPLE ID	WT (mg)	%H ₂ O	% RESIDUE	% CARBO-HYDRATE	OXALATE	
					425 540 C	631 750 C
70515	9.8270	6.9401	9.4739	68.5359	6.0141	6.4109
70518	9.9530	5.5059	5.3753	76.3485	5.0638	2.9941
70519	10.4880	3.4134	8.1998	68.3159	5.9497	6.4836
70523	9.5570	3.2855	6.8013	66.6212	5.9642	7.3663
70526	9.8250	5.8321	8.1629	74.7279	4.8651	3.1959
AVERAGE (X)	9.9300	4.9954	7.6026	70.9099	5.5714	5.2902
STD DEV	0.3437	1.5944	1.5611	4.3276	0.5990	2.0401
N = 5						

TABLE 1E
THERMOGRAVIMETRIC ANALYSIS
OF RECYCLED CARRISYN
40 MINS. AGITATION

SAMPLE ID	WT (mg)	%H ₂ O	% RESIDUE	% CARBO-HYDRATE	OXALATE	
					425 540 C	631 750 C
70528	9.6900	5.7069	6.1197	77.7195	4.7678	2.633
70529	9.2960	7.0998	4.9699	75.8714	4.5611	3.001
70535	9.7970	6.5428	4.3483	79.3609	4.2462	1.194
AVERAGE (X)	9.5943	6.4498	5.1460	77.6506	4.5250	2.182
STD DEV	0.2638	0.7011	0.8987	1.7457	0.2627	0.922
N = 3						

The tables reveals that:

1. Carrisyn manufactured from a starting gel having a pH in excess of 4.56 demonstrated average total carbohydrate and residue (ash) contents of

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52.2% and 18.1% respectively.

2. With adjusted pH of starting gel from about 4.56 to 3.60, the carbohydrate and ash contents were 71.4% and 8.55% respectively. This is an improvement of 36.8% and 52.8% for each measured parameter.
3. However, with a pH of 3.20, the average carbohydrate content jumped to 79.3% while ash content decreased to 3.86% signifying a carbohydrate increase of about 51.2% and ash decrease of 78.7%.
4. Preliminary data demonstrate that Carrisyn batches with increased carbohydrate content and lower ash value demonstrated better antiviral activity; it is therefore recommended that processing be conducted at a pH between 3.00 and 3.50.

TABLE 2

Effects of pH on Quality of Carrisyn - (TGA METHOD)

pH	2.45	3.35	4.07	4.56	5.09	6.16	7.00	8.68
Moisture	5.6762	6.2433	6.9750	8.4035	8.5905	8.8431	8.9819	9.7702
Acemannan	73.925	69.930	46.929	38.651	37.649	36.796	35.593	36.316
Carbon Skeleton	17.812	15.455	12.840	12.846	12.969	12.508	11.485	10.769
Calcium	0.4990	4.3160	9.7965	10.670	10.251	10.347	11.725	10.148
Oxalate								
Ash	1.9644	3.9494	13.811	15.999	16.847	17.908	17.192	17.198
Yield %	0.12	0.14	0.22	0.26	0.27	0.25	0.26	0.24

Carrisyn (total solid) yield, calcium oxalate, ash content and moisture appear to increase with increased pH up to 4 - 5. But acemannan and the corresponding carbon skeleton increase with decreased pH. Therefore a pH greater than 4.0 is not recommended if ash and calcium oxalate contents must be

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decreased. A yield of Carrisyn greater than 0.2% must be analyzed for calcium oxalate contamination.

By treating Aloe vera gel with a suitable acid, e.g., a dilute mineral acid, preferably hydrochloric acid, to adjust the pH of the gel to between 3.0 to 3.5 followed by ethanol extraction effected a significant reduction in the amount of oxalates. By this step, both calcium oxalate and ash content may be reduced by more than 80%. The treatment also concentrated the amount of the active Carrisyn without degrading the polymer as demonstrated by physicochemical analytical methods.

Two batches of Carrisyn were manufactured as described in Example 1, but with the additional step of acidification to pH 3.6 prior to ethanol precipitation. In Lot # 70315, concentrated nitric acid (61 ml) was added to 20 gallons of homogenized aloe vera gel. In Lot # 70321, concentrated hydrochloric acid (171 ml) was added to 45 gallons of homogenized aloe vera gel. The yields were 55.6g (0.07%) and 186.6g (0.11%), respectively. Chemical analysis by IR and TGA showed both batches to have similar quality and reduced calcium oxalate. The analyses of these two batches are included in Table 1B.

It should become apparent to one skilled in the art, that nitric acid can be used solely for acidification, as well as hydrochloric acid or any other suitable acids.

CHARACTERIZATION OF CARRISYN

Using pharmaceutical screening techniques a polysaccharide extracted from Aloe vera has now been found to be the active chemical substance in Aloe vera. This polysaccharide will be hereinafter referred to as acemannan. Acemannan is an ordered linear polymer of substantially acetylated mannose monomers. Other ingredients, such as proteins, organic acids, anthraquinones, vitamins and amino acids make up less than one percent of Carrisyn™. The concentration of

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Carrisyn™ in Aloe vera has been found to be approximately 0.05 to 0.3 weight-percent of the Aloe vera juice. The yield or concentration of Carrisyn™ in the leaves depends on leaf maturity.

The pharmacological data that evidences that acemannan is the active chemical substance in Aloe vera can be summarized as follows:

1. The dose-response of acemannan was the same as Aloe vera juice with an equivalent amount of acemannan.

2. Acemannan was effective in the ulceroprotection model by different routes of administration namely intravenously, intraperitoneally and orally.

3. Acemannan accounted for 100 percent of the effects in both pharmacological models.

4. The chemical substance, glucomannan a substance similar to acemannan from a completely different source, the Konjac plant, provided some pharmacological response.

Carrisyn™ has been shown in laboratory studies to increase up to 300% in 48 hours the replication of fibroblasts in tissue culture which are known to be responsible for healing burns, ulcers and other wounds of the skin and of the gastrointestinal lining.

Carrisyn™ has also been shown to increase DNA synthesis in the nucleus of fibroblasts. The increase in DNA synthesis in turn increases the rate of metabolic activity and cell replication which are fundamental steps to the healing process.

Carrisyn™ has been shown in controlled studies to increase the rate of healing in animals.

Carrisyn™ has also been shown to be an effective treatment for gastric ulcers in animal studies. Over a three year period laboratory rats, the stomachs of which react similar to that of humans, were tested. Carrisyn™ was found to be equivalent to or superior to current medications used for the treatment of gastric ulcers. Most such products act to inhibit hydrochloric acid in the stomach. Carrisyn™ works on a

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different principle and does not alter the natural flow of digestive acids.

As noted above, Carrisyn™ can be precipitated out of liquidified Aloe vera gel by the addition of a water soluble, lower aliphatic polar solvent, preferably ethyl alcohol. Carrisyn™ powder can then be prepared by lyophilization and optionally the lyophilization product can be ground into a powder with a grinding apparatus such as a Moulinex coffee-grinder (available from Dillard's, Dallas, Texas). Carrisyn™ powder is highly electrostatic and is an off-white to pinkish-purplish amorphous powder depending on the oxidization state of any anthraquinone contaminant. Carrisyn™ is stabilized and becomes substantially non-degradable by freeze drying or lyophilization which removes the water which causes hydrolysis. Freeze dried Aloe vera gel with a given amount of Carrisyn™ has maintained its effectiveness for two years. It is believed that Carrisyn™ in freeze dried form will be stable for up to ten years.

Heat and time have been found to be important factors in the production of powdered Carrisyn™. Heat can aid in the hydrolysis or degradation of Carrisyn™ and the longer it takes to process out the Carrisyn™ at a given temperature, the more is the degradation. Accordingly, it is preferred that the process which allows for the quickest extraction of Carrisyn™ from whole Aloe vera leaves be used when high molecular weight Carrisyn™ powder is desired and it is preferred that the process which allows for the slowest extraction of Carrisyn™ from whole Aloe vera leaves be used when low molecular weight Carrisyn™ powder is desired.

Rehydration of Carrisyn™ powder at a weight/volume concentration of 0.2 to 1 percent resulted in the reformation of a viscous "gel" like fresh Aloe vera. The gel-like consistency which returns upon rehydration of Carrisyn™ powder is indicative of the high molecular weight polysaccharidic nature of Carrisyn™. Generally, as polysaccharides are degraded or hydrolyzed their viscosity is lowered. Accordingly, the

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viscosity of rehydrated Carrisyn™ powder gives a good indication of quality and may be used as a parameter for quality assurance.

Carrisyn™ produced according to the process of the present invention may be characterized as a substantially non-degradable lyophilized linear polymer of substantially acetylated mannose monomers, preferably bonded together by 1 - 4) bonds.

Various modifications of the disclosed compositions of the invention, as well as alternative modifications, variations and equivalents will become apparent to persons skilled in the art upon reading the above general description. The following Examples (Examples 5 - 8) were conducted in order to further characterize and identify Carrisyn™. The following examples are illustrative only and are not intended to limit the scope of the appended claims, which cover any such modifications, equivalents or variations.

EXAMPLE 5

Isolation, Purification And Characterization Of Carrisyn™

A. Isolation of Carrisyn™

An aloe leaf was washed, sliced open and filleted. The clean inner gel was retained while the green rind and latex materials were discarded. The filleted material was homogenized and extensively filtered with a Finisher Model 75 (FMC, Chicago, IL), to remove most of the pulp. The clear, viscous gel was acidified to a pH of approximately 3.20 with dilute HCl to solubilize the oxalates and lactates of calcium and magnesium that are usually present to their corresponding water soluble acids. The acid treated gel was then extracted for 4-5 hours with four volumes of 95% ethanol at ambient temperature. Floating fibers were removed, then the alcohol/-water mixture was siphoned off while the solid precipitate was collected by centrifugation. Most alcohol/water soluble

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substances such as organic acids, oligosaccharides, monosugars, anthraquinones and inorganic salts were eliminated in the process. The solid was then washed with fresh alcohol, centrifuged, freeze dried, and ground to a white amorphous powder.

B. Purification

Carrisyn at this stage is generally contaminated with proteins, monosugars, oligosaccharides and inorganic salts. These contaminants do not affect the bioactivity of the product hence further purification is not necessary for manufactured bulk material. However, as a further step in the characterization of Carrisyn, more purification steps were added. The above-mentioned contaminants were removed by redissolving the bulk powder in phosphate buffer and treating it with non-specific protease (Sigma Lot #5147) followed by extensive dialysis. The non-filterable product, which is mainly acetylated polymannose, was freeze dried and characterized using a number of analytical methods which include IR spectroscopy, thermogravimetric analysis (TGA), HPLC, GLC and GC/MS as depicted in Figure 11.

For the standardization of Aloe vera products HPLC and GC are recommended. The gas liquid chromatographic (GLC) method is used to separate and quantitate mannose in aloe vera gel extract. In order to prevent the spiking of products with mannose to make a product appear to have a higher Aloe gel content, a dialysis step may be added (MW cutoff 12,000 - 14,000). In addition, spiking with locust bean gum, guar gum or glucomannan could be detected by a high galactose or glucose to mannose ratio.

C. Molecular Weight Determination

Introduction:

Acemannan (AM) is a plant extract polysaccharide. It is polydispersed which means that it is comprised of more than one molecular weight size.

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Objective:

The object of the study is to determine the molecular weight distribution of acemannan by size exclusion chromatography.

Reagents and Chemicals:

0.05% sodium azide.

Standards:

0.2% (w/v) of each standard, Pullulan 853K, 100K and 12.1K daltons in 0.05% sodium azide. (Shodex P-82, Showa Denko, K.K.)

Instrumentation:

High Performance Liquid Chromatograph, model 590 (Waters Associates, Milford, MA)

Differential refractometer, model 1770 (Bio-rad)

Integrator SP 4290 (Spectra-Physics)

Sample Preparation:

-Weigh 20 mg of Carrisyn into a glass vial (105x25mm) with a teflon lined cap.

-Add 10 mL of 0.05% sodium azide and dissolve by shaking (4 hours) in a Junior orbit shaker (Lab-Line Instruments, Mel Rose Park, IL).

-Filter through a 1.2 um membrane, (Acrodisc(R), Gelman Sciences). Save the filtrate for an injection into the HPLC.

High Performance Liquid Chromatographic (HPLC)Conditions

Column: Spherogel TSK 5000 PWHR (Beckman Instruments)

Detector: Differential refractometer (Bio-rad)

Mobile Phase: 0.05% sodium azide

Flow rate: 1 mL/min at 40°C

Vol. inj.: 50 uL

Results:

Acemannan is a polydisersed polysaccharide with at least 73% of the polymer greater than 10,000 daltons as determined by size exclusion chromatography (SEC) and pullulan polysaccharide as the standard. Figure 12, represents the SEC

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chromatogram of pullulan standard of known molecular weights; 853K, 100K and 12.1K daltons respectively. The corresponding chromatogram of Carrisyn is shown in Figure 13. The chromatogram highlights three (3) main peak fractions, labeled A, B, and C. Peak A represents the acemannan fraction greater than 100,000 daltons and peak B is the fraction greater than 10,000 but less than 100,000 daltons. Peak C represents lower molecular weight constituents. The sum of peaks A and B constitute the active fraction.

D. Infrared (IR) Spectroscopic Analysis

Introduction:

The functional groups of the acemannan, the active product of Carrisyn™, absorb at characteristic infrared frequencies. Infrared spectroscopy therefore is an important method to characterize this material.

Objective:

The objective of the study is to further characterize Carrisyn by an infrared spectroscopic method.

Reagents and Chemicals:

Infrared grade potassium bromide (KBr) powder (Mallinckrodt Inc., Paris, Kentucky).

Instrumentation:

IBM Fourier Transform infrared (Ft-Ir) spectrometer model #32 equipped with an IBM 9000 computer and a printer/plotter.

Sample Preparation:

-Carrisyn is preground to fine powder by using Wiley Mill (Thomas Scientific Co.) and a screen which allows particles smaller than 60 mesh size.

-A 5 mg preground sample is mixed with 495 mg of dry KBr to a total of 500 mg mixture.

-The mixture is reground by hand using agate mortar and pestle to a homogeneous material.

-A representative sample (80-100mg) is pressed into a transparent disc using a hydraulic jack (Walker^R) at a

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pressure of 40,000 psi.

-The disc is scanned from 4000 cm^{-1} to 400 cm^{-1} . A multiple scan (30 scans) is performed with a 4 cm^{-1} resolution to improve signal to noise ratio.

Results:

Analysis of the spectrum depicted in Figure 14, highlights the following characteristic absorption peak frequencies (cm^{-1}):

Wavenumber(cm^{-1})	Assignment
1066.8	C-O stretch of pyranose ring structure
3422.1	O-H stretch
1250.0	C-O-C stretch (acetyl group)
1740.0	C=O stretch (acetyl)
1377.3	C-H bending
1649.3	C=O stretch (Amide I)
530.5	Internal rotation modes etc.
599.9	Internal rotation modes etc.
2928.3	C-H stretch
1541.3	N-H deformation (Amide II)
806.3	-----
897.0	Axial H on ring at C_1
773.6	ring breathing

It is noted that the spectrum of Carrisyn demonstrates a strong absorption due to O-H stretching about 3422 cm^{-1} . The carbonyl and C-O-C stretches of acetyl group are located near 1740 and 1250 cm^{-1} respectively. The strong single band of C-O-C system is indicative of O-acetyl group bonded equatorially to the monomer unit. The amide carbonyl stretch (amide I) at about 1649 cm^{-1} superimposed on the moisture absorption peak and the N-H deformation (amide II) about 1541

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cm⁻¹ are due to protein/proteoglycan impurities.

The absorption bands between 960-730 cm⁻¹ can be correlated with certain stereochemical features of acemannan. For example, the absence of a band near 844 cm⁻¹ due to equatorial C₁-H and the presence of axial C₁-H near 897 cm⁻¹, demonstrate that the acemannan is B-linked. The ring vibration (955 cm⁻¹ shoulder) and the ring breathing located about 773 cm⁻¹ indicate a D-mannan with beta-glycosidic linkage.

Table 3 below depicts the peak absorption frequencies with corresponding absorbances arranged according to intensity:

TABLE 3

PEAK ABSORBANCE FREQUENCIES FOR
NON-PROTEASE TREATED CARRISYN

Peak #	Peak	Peak Start	Peak End	Abs
1	1066.8	1201.8	1049.4	.546
2	3422.1	3433.7	3410.6	.474
3	1250.0	1348.4	1201.8	.439
4	1740.0	1821.0	1699.5	.433
5	1377.3	1410.1	1350.3	.343
6	1649.3	1697.6	1585.7	.325
7	530.5	582.6	493.8	.313
8	599.9	634.7	586.4	.282
9	2928.3	2995.8	2899.4	.242
10	1541.3	1581.8	1527.8	.226
11	806.3	835.3	789.0	.224
12	897.0	920.2	843.0	.210
13	773.6	787.1	750.4	.207

E. Infrared Spectroscopy of Protease Treated Carrisyn

The infrared spectrum of bulk Carrisyn revealed traces of protein or proteoglycan impurities as demonstrated by the presence of amide I and amide II peaks. Further purification of the bulk material with non-specific protease which hydrolyzes the protein or proteoglycan followed by extensive dialysis, eliminated these impurities.

The spectrum of the protease treated sample is shown in Figure 15. Analysis of the spectrum shows some differences when compared to the non-protease treated Carrisyn of Figure

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14. For example, the amide I and II peaks located about 1649 and 1541 cm^{-1} in Figure 14 are absent from the protease treated sample. In addition, the moisture absorption peak located about 1639.7 is clearly resolved. Table 4 represents the peak absorption frequencies with corresponding absorbances arranged according to intensities for protease treated Carrisyn as follows:

TABLE 4

PEAK ABSORBANCE FREQUENCIES FOR
PROTEASE TREATED CARRISYN

Peak #	Peak	Peak Start	Peak End	Abs
1	1064.8	1093.8	1049.4	.969
2	1032.0	1045.5	970.3	.966
3	3422.1	3435.6	3404.8	.814
4	1248.1	1348.4	1201.8	.729
5	1740.0	1838.4	1695.6	.679
6	1377.3	1406.3	1350.3	.593
7	530.5	545.9	493.8	.575
8	598.0	677.1	586.4	.532
9	1639.7	1695.6	1554.8	.491
10	2928.3	2990.0	2903.2	.428
11	897.0	918.2	887.4	.409
12	806.3	835.3	790.9	.409
13	771.6	789.0	748.5	.387

The absorbance values are not comparable in intensities with the non-protease treated Carrisyn of Figure 14 since it is qualitative and no attempt was made to use the same concentration in making the sample discs.

On the basis of infrared spectroscopy alone, Carrisyn is a polysaccharide of essentially B-linked D-mannose with O-acetyl group side chains. The presence of N-acetyl groups may be due to protein/proteoglycan impurities.

F. Thermogravimetric Analysis (TGA) of Carrisyn

Introduction:

Thermogravimetric analysis (TGA) is an important analytical method for studying polymers. The mass loss on decomposition of a polymer as a result of temperature change is characteristic of that polymer. Moreover, TGA aids in the determination of moisture (H_2O) and ash content of powdered

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substances.

Objective:

It is the objective of this study to use TGA to further characterize Carrisyn.

Reagents and Chemicals:

None.

Instrumentation:

1. Mettler Thermoanalyzer, TA 3500 system featuring a TC 10A Evaluation and Control computer, as well as a TG 50 furnace with M3-03 microbalance.
2. Printer/plotter model MP3 (Print Swiss Matrix)
3. IBM PC for file storage.

Sample Preparation and Analysis:

-A 10 mg sample in a 70 mL alumina crucible is weighed in a microbalance accurate to ± 1 ug.

-The crucible with its contents is heated in the TG 50 furnace at a temperature program rate of 20°C/min. This rate of heating is standard and adequate for the material to be analyzed.

-Heating is performed in a nitrogen gas atmosphere from 25°C to 600°C and then from 601°C to 780°C in air (oxidant) atmosphere.

-The temperature is held at this final temperature for an additional (2) two minutes. (The maximum temperature of 780°C is chosen because both organic and inorganic substances decompose before this temperature is reached.)

Results:

The decomposition profile of Carrisyn is characteristic. Figure 16 depicts the real time plot and the corresponding first derivative plot of the decomposition profile of Carrisyn. The acemannan (the active substance of Carrisyn) decomposition pattern is different from those of other polysaccharides. For example, under identical operating conditions, the temperature at which acemannan exhibits major decomposition is different from those of cellulose, dextran or amylan. The weight loss associated with acemannan is located

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between 200°C to 630°C. The bulk of Carrisyn's decomposition occurs between 200°C to 540°C and 600°C to 630°C. The aceman-nan fraction is determined by the contributions of these two temperature regions, while ash and moisture together contribute about 10% by weight. Table 5 below represents a replicate analysis of Carrisyn with TGA.

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TABLE 5

TGA ANALYSIS OF CARRISYN

Sample #	Weight (mg)	H ₂ O %	Ash* %	AM %	CaOx %
1	9.4390	7.3631	3.8034	82.2119	2.5532
2	9.5860	7.3649	4.0163	81.8381	2.7540
3	9.5210	7.4992	4.2012	81.7568	2.9514
4	9.4530	7.7118	4.1786	81.2010	2.9620
5	9.4090	7.7691	3.8793	81.6072	2.9546
6	9.5960	6.8674	3.9912	82.0969	3.1576
7	9.4740	7.6209	4.1904	81.6865	2.8605
8	9.5730	7.3436	4.2411	81.0719	3.3218
9	9.5550	7.9435	4.1130	80.9652	3.2444
10	9.5220	7.7715	4.3373	80.8023	3.1611
Ave.** (X)	9.5128	7.5255	4.0952	81.5238	2.9920
Std Dev.	0.0626	0.3087	0.1688	0.4864	0.2353

*The ash consists of the oxides of Ca (1.51), Si (0.1), Na (.55), Mg (.37), Fe (.02) and Al (.00).

**Average value will not total 100% because the process does not account for peaks less than 2% of the base peak.

This method of analysis is qualitative as well as semi-quantitative. For example, the percent acemannan may be determined as follows:

$$1. \quad \text{Percent (\%)} \text{ AM} = \frac{\text{mg (200-630}^\circ\text{C)} \times 100}{\text{mg (sample)}}$$

or

$$2. \quad \text{Percent (\%)} \text{ AM} = \frac{\text{mg (200-630}^\circ\text{C)} \text{ sample} \times 100}{\text{mg (200-630}^\circ\text{C)} \text{ ref. std.}}$$

The moisture and ash contents are important parameters in powdered drug substances. These parameters are easily determined using the Thermogravimetric method.

G. Constituent Sugar Determination of Carrisyn
By Acid Hydrolysis and High
Performance Liquid Chromatography (HPLC)

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Introduction:

Polysaccharides are hydrolyzed to their constituent monomers by acid or enzymatic hydrolysis. A 2 M trifluoroacetic acid (TFA) is chosen for the hydrolysis because it is strong enough to hydrolyze the glycosidic bonds but unlike sulfuric acid it is gentle enough that the monomeric sugar residues are not destroyed.

Objective:

The objective is to determine the monomeric sugar constituents of the acemannan fraction of Carrisyn.

Reagents and Chemicals:

1. 2 M TFA containing 0.5 mg/mL inositol as the internal standard.
2. Isopropanol

Instrumentation:

1. Hewlett Packard HPLC model 1084B equipped with an autosampler HP79842A and auto injector HP79841A.
2. Bio-rad differential refractometer model 1770.

Sample Preparation:

-Weigh accurately 2.0 mg of Carrisyn on weighing paper and transfer quantitatively to a (13x100 mm) disposable culture tube with a teflon lined cap.

-Add one milliliter (mL) of 2M TFA containing 0.5 mg/mL inositol as an internal standard.

-Place the tube in a heating block (Hycel Thermal Block) at 120°C for approximately 1 hour.

-Evaporate the hydrolysate mixture to dryness with dry air.

-Redisperse the solid in 1 mL of isopropanol and evaporate to dryness with dry air.

-Dissolve the solid in 1 mL of deionized water in preparation for injection into the HPLC column.

HPLC Conditions:

Column: Aminex Carbohydrate HPX-87P

(Bio-rad Labs., Richmond, CA)

Mobile phase: Deionized water at 80°C

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Flow rate: 0.6 mL/min

Oven temp: 40°C

Chart speed: 0.2 cm/min

Detector: Refractive index (Bio-rad #1770)

Attn: 2

Results:

Figure 17 represents the chromatogram of a standard mixture, comprising glucose, galactose, mannose 1 mg/ml each and 0.5 mg/ml inositol as the internal standard. Figure 18, represents the chromatogram of Carrisyn under identical conditions. It is noted that mannose is the major component of the polymer signifying that the polysaccharide is essentially composed of mannose sugar units.

H. Gas Liquid Chromatographic Separation and Quantitation of CarrisynIntroduction:

The polysaccharide of Carrisyn is mainly neutral in nature. Neutral polysaccharides are better analyzed by Gas Liquid Chromatography (GLC) as their glycol acetates.

Objective:

The objective of this study is to further characterize Carrisyn by separating and quantitating the sugar monomers as their glycol acetates. By this procedure, the glycol acetates are separated with a capillary column and detected by a flame ionization method.

Reagents and Chemicals:

1. A 2M TFA containing 0.2 mg/mL inositol as the internal standard
2. Isopropanol
3. 1M NH_4OH containing 10 mg/mL sodium borodeuteride (NaB^2H_4) or sodium borohydride (NaBH_4)
4. Glacial acetic acid
5. Methanol
6. Pyridine
7. Acetic anhydride
8. Toluene and dichloromethane

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9. Gas chromatography (G.C.) grade acetone

Instrumentation:

1. Gas chromatograph Vista 6000 (Varian Instrument Group, Palo Alto, CA)
2. Integrator SP4290 (Spectra-Physics, San Jose, CA)

Sample Preparation:

-Weigh accurately 2 mg of sample and transfer into a culture tube (13x100 mm) with a teflon lined cap.

Hydrolysis:

-Add 500 uL trifluoroacetic acid TFA-(2M, containing 200 mg/mL inositol), heat at 120°C (Hycel Thermal Block) for about one hour.

-Remove TFA (water bath at 40°C) under flow of air and wash residue with isopropanol to remove residual acid.

Reduction:

-Dissolve residue in 1 M NH_4OH (300 mL) containing 10 mg/mL sodium borodeuteride (NaB^2H_4) leave at room temperature for 1 hour.

-Destroy excess reductant with glacial acetic acid (few drops until no effervescence), remove solvents under air and wash residue with methanol containing 10% acetic acid (3x300 uL) and finally methanol (3x300 uL).

Acetylation:

-To the dry residue add 200 uL pyridine and 200 uL acetic anhydride, heat for 20 min at 120°C.

-To the cooled solution add toluene 500 uL, remove solvent under air, dissolve residue in water and extract into dichloromethane.

-Transfer the organic phase to a clean tube, remove the organic solvent under air and dissolve residue in acetone (100 uL), prior to GLC.

Gas Liquid Chromatographic Conditions (GLC)

Column: SP 2330, 15 M or 30M, 0.25 mm I.D., 0.25 um liquid phase thickness (Supelco, Inc.)

Carrier gas: Helium

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Oven temp.: 235°C (isothermal)

Inj. temp: 250°C

Inj. vol.: 0.5 - 1 uL (split)

Results:

Figure 19, represents the GLC chromatogram of a standard mixture of rhamnose, fucose, arabinose, xylose, mannose, galactose, glucose, and inositol as their glycol acetates. The chromatogram of Carrisyn is shown in Figure 20. The major peak in the figure, corresponds to mannose. There are traces of galactose, glucose, arabinose, xylose, and fucose. These are from cell wall contaminants. The last peak is inositol acetate which is the internal standard.

Since mannitol acetate is the major sugar peak in the hydrolysed and acetylated material, Carrisyn is essentially a polymannose polysaccharide.

The mannitol acetate peak can be quantitated with the help of inositol as the internal standard. Figure 21, represents a standard curve generated by plotting mannose/inositol area ratio against the known weights of acemannan. With the standard curve, the amount of acemannan in a given sample of Carrisyn can be calculated.

I. Gas Chromatography/Mass Spectrometry
and Glycosidic Linkage Analysis of Carrisyn

Introduction:

Polysaccharides are characterized by determining the manner in which the sugar units are linked to one another. The chemical, physical, and biological activities of polysaccharides depend on where the sugars are linked in the five possible positions for the hexoses.

Objective:

It is the purpose of this study to determine the manner in which the sugar units of Carrisyn polysaccharide are linked.

Reagents and Chemicals:

1. Dry dimethylsulfoxide (DMSO)
2. Sodium dimethylsulphanyl anion (2M)

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3. Methyl iodide
4. 2M TFA
5. Isopropanol
6. NH_4OH with 50% methanol, 10 mg/mL (NaB^2H_4)
7. Glacial acetic acid
8. 10% acetic acid in methanol
9. Acetic anhydride
10. Sodium bicarbonate
11. Dichloromethane
12. GC grade acetone

Instrumentation:

Gas chromatograph/mass spectrometer MSD (HP 5970).

GLC/MS and Glycosidic Linkage Analysis

The overall characterization of Carrisyn includes the linkage analysis of the polymer. Carrisyn is permethylated by the method described by Darvil et al., (Plant Physiology, 62: 418 - 422 (1978)), the disclosure of which is hereby specifically incorporated herein by reference, hydrolyzed to the monomers and converted to methylated alditol acetates. The volatile derivatives are analyzed on Hewlett-Packard GC/MS system on a Supelco SP 2330 capillary column (30 m x .25 mm id). Fragmentation of the methylated glycitol acetate of the monomers is achieved by electron impact (EI) method. The total ion chromatogram (TIC) of the derivatized Carrisyn as the partially methylated and partially acetylated glycitol is shown in Figure 22. The mass spectrum of the partially acetylated mannitol is demonstrated in Figure 23.

The complete procedure is as follows:

- (1) The sample (1 mg) is placed in a tube with a teflon lined cap and dried in a vacuum oven at 50°C.
- (2) To the sample is added dry DMSO (250 uL) and the sample stirred (teflon stir-bar) until it

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dissolves (sonication may help). Purge tube with argon (or nitrogen).

- (3) To the polysaccharide (in DMSO) is added 250 uL of sodium dimethylsulphinyll anion (2M) and left stirring for a minimum of 4 hours (reaction should be performed under argon). It may be more convenient to leave the sample overnight in the anion solution.
- (4) The anion solution is cooled in ice, methyl iodide 200 uL is slowly added and the solution stirred for about 1 hour.
- (5) To the solution, add about 2.5 mL water and remove excess methyl iodide with argon.
- (6) Transfer the mixture to dialysis tubing, 13.3 x 2.1 cm in diameter (M.W. cutoff 12,000 - 14,000, Spectrum Medical, Inc. Los Angeles, CA).
- (7) Concentrate the non-dialysable fraction under a flow of air (50°C).
- (8) To the dry residue add 250 uL of 2M TFA and heat for 1 hour at 120°C.
- (9) Remove TFA under air, wash residue with isopropanol (2 x 250 uL).
- (10) Dissolve residue in aq. NH_4OH 50% methanol (250 uL) containing 10 mg/mL NaB^2H_4 leave 1 hour at room temperature.
- (11) Destroy excess reductant with glacial acetic acid (few drops), concentrate to dryness and wash residue with 10% acetic acid in methanol (3 x 250 uL).
- (12) To the dry residue, add acetic anhydride (100 uL) and heat sample at 120°C for 3 hours. To the cooled sample add water (about 1.5 mL) and then sodium bicarbonate until effervescence stops. Extract the derivative into dichloromethane.

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(13) Concentrate organic phase to dryness, dissolve residue in 100 uL acetone prior to GLC and GLC/MS.

GLC conditions:

Column: . SP 2330 fused silica column
(30 m x 0.25 mm).

Temp.: 3 min at 170°C followed by 4°C/min to 240°C hold for 10 min.

Detector: FID

GLC/MS conditions:

Column: SP 2330 (30m x 0.25 mm)

Temp.: 2 min at 80°C, increased to 170°C at 30°C/min then 4°/min to 240°C hold for 10 min

MS: Hewlett Packard MSD.

Determination of the Major Glycosidic Linkages of Samples by Methylation Analysis

In this procedure all the free hydroxyl groups of the polysaccharide are converted to methyl ethers (step. 4). The methylated polysaccharide is hydrolyzed to the constituent monosaccharides (step 8), converted to the methylated alditols (step 9) and acetylated (step 10). These volatile derivatives are analyzed by GLC/MS (step 13) and from the fragmentation patterns, the positions of the 0-methyl and 0-acetyl groups on the alditols are determined.

Illustration

Consider a glycosyl (mannosyl) residue linked through positions 1 and 4 (ie a (1 - 4) - linked mannan). The hydroxyl groups at positions 2, 3 and 6 are free and can be converted to methyl ethers. On hydrolysis and reduction the hydroxyl groups at positions 1, 4 and 5 are exposed and can be acetylated to yield the derivative 1, 4, 5-tri-O-acetyl-2, 3, 6-tri-O-methylhexitol (Figure 24). When examined by mass spectrometric technique, the dominant primary fragment ions

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are produced by cleavage between adjacent O-methyl groups or between O-methyl and O-acetyl groups. The secondary fragment ions are produced by loss of acetic acid, methanol, etc. (Figure 25). Figures 22 and 23 demonstrate the total ion chromatogram of Carrisyn and the mass spectrum of 4-linked mannan as the partially acetylated mannitol respectively.

Based on the total ion chromatogram (TIC) of the Carrisyn and the linkage analysis, it is deduced that Carrisyn is mainly a (1 - 4)-linked linear polymer of mannose.

EXAMPLE 6

A 32 year old patient, was presented with a history of ulcerative colitis for "many years". During an active episode, she had been unresponsive to a daily regimen of 40mg Prednisone, 3 grams Asulfidine, 50 mg 6-mercaptopurine, and Flagyl. She continued to have a painful abdomen and 4-8 bloody bowel movements per day. She was placed on hyperalimentation. Endoscopic findings revealed severe ascending colon ulcerations with mild hepatic to transverse ulcerations. The patient was placed on 50 mg of Carrisyn™ q.i.d. in addition to her other medications and sent home. In one week, her symptoms were virtually gone. The abdomen was mildly tender and endoscopy revealed a healed and mildly congested mucosa. The patient was slowly taken off other medications and the clinical picture continued to improve. The patient is currently maintained on Carrisyn™ as the sole medication at this time. Physical exam and symptoms are recorded as totally normal.

Five additional cases with similar responses to ulcerative colitis and Crohn's disease have been seen. One patient ran out of Carrisyn™ capsules. In four weeks, mild symptoms began to recur (there was increased stool with mild abdominal discomfort), and she returned for a supply of medication. In three days she was back to total normal bowel symptomatology.

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EXAMPLE 7

A number of AIDS patients have received prolonged high doses of Carrisyn™ without toxicity or side-effects. A rise in T-4 and T-8 lymphocyte ratios and an increase in absolute T-4 counts was seen in these AIDS patients with a reduction and elimination of clinical symptoms, as well as a reduction in opportunistic infections. It is suggested that Carrisyn™ had an anti-viral or immune modulation effect in patients.

A stimulation to the lymphocytes of these patients has been observed which suggests that Carrisyn™ may be involved in immune modulation.

EXAMPLE 8The Effect of Alcohol Concentration on Carrisyn™ Yield

PROCEDURE:

Hilltop leaves (15.9 lbs.) were washed, filleted and ground in a Waring blender, then filtered through eight layers of cotton cloth. The gel was then transferred to four 11 quart, stainless steel pans, and cold USP grade ethyl alcohol was added to each in two, three, four and five to one ratios by volume. The amounts can be summarized as follows:

<u>Ratio (ethanol:aloe gel)</u>	<u>Amt. of Gel</u>	<u>Amt. of Ethyl Alcohol</u>
2:1	500 ml	1000 ml
3:1	500 ml	1500 ml
4:1	1670 ml	6680 ml
5:1	500 ml	2500 ml

The precipitates were allowed to settle out for four hours, then the remaining alcohol-gel solutions were carefully decanted and saved in separate containers. The precipitates were centrifuged for 10 minutes at 2800 rpm using a IEC Centra-7 centrifuge, washed with alcohol, then centrifuged again under the same conditions. The pellets were transferred to 600 ml jars, frozen in liquid nitrogen, and lyophilized overnight.

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An additional volume of alcohol was added to the supernatant from the 2:1 ratio and allowed to settle overnight at room temperature. The remaining supernatants were also left at room temperature and allowed to settle out overnight.

The following day, the precipitates were collected from the supernates as previously described with the exception of the pellet from the 2:1 ratio that had been precipitated with an additional volume of alcohol. In this case, approximately 5-10 ml of water was added as the pellet was transferred to the lyophilization jar.

RESULTS

The results of the initial, four hour alcohol precipitations can be summarized as follows:

<u>Ratio (ethanol:aloe gel)</u>	<u>Yield(g)</u>	<u>%Yield(g. Carrisyn™/g. gel)</u>
2:1	.0518	.010
3:1	.3847	.077
4:1	1.945	.116
5:1	.6675	.134

After addition of another volume of ethyl alcohol, the 2:1 supernate produced another 178 mg of Carrisyn™. Just by settling overnight, the 3:1 and 4:1 ratio supernates yielded another 89 and 105 mg, respectively. The 5:1 ratio yielded only negligible precipitation after the initial isolation, and was thus not reharvested.

In the case of the second precipitation from the 2:1 ratio (3:1), 5-10 ml of water were used to rinse out the centrifuge bucket before lyophilization. This produced a white, fluffy Carrisyn™ of low density that differed greatly from the denser, gray colored Carrisyn™ samples that the other samples produced.

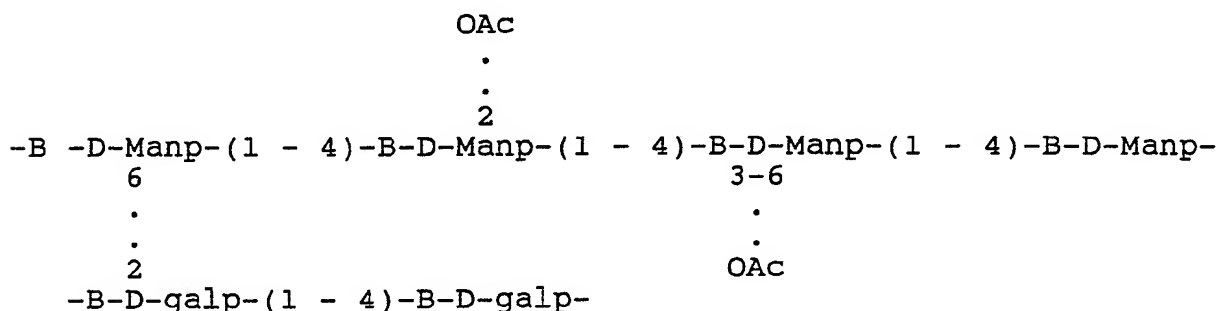
SUMMARY

Carrisyn™ is a purified white amorphous powder extracted from Aloe vera mucilage. The polymer is essentially made up of linear B (1 - 4)-D-mannosyl units. It is a long

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chain polymer interspersed randomly with acetyl groups linked to the polymer through an oxygen atom., The degree of acetylation is approximately 0.8 acetyl group/monomer as determined by the alkaline hydroxamate method (Hestrin, S.: J. Biol. Chem., 180: 240 (1949)) the disclosure of which is hereby specifically incorporated herein by reference. Neutral sugars linkage analysis indicates that attached to the chain, probably through an (2 - 6) linkage, is a D-galactopyranose residue in the ratio of approximately one for every seventy sugars. The ratio of mannose to galactose of 20:1 indicates that galactose units are also linked together, primarily by B (1-4) glycosidic bonds.

The chemical structure developed by utilizing modern techniques of polysaccharide characterization is as follows:



Carrisyn™ is polydispersed with at least 70% of the material having a molecular weight greater than 10,000 daltons.

Physical and Chemical Properties:

(a) Solubility

Carrisyn is a white to off-white amorphous powder which dissolves slowly in water to a highly viscous colloidal solution (0.4% w/v). With vigorous shaking for several hours, a 1% (w/v) thick gel may be obtained. Carrisyn is practically insoluble in common organic solvents including propylene glycol. However, Carrisyn dissolves in 20% water and 80% propylene glycol to a smooth thick gel which remains stable indefinitely.

(b) pH:

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A 0.2% (w/v) solution of Carrisyn in water has a pH of approximately $6.31 \pm .33$.

(c) Optical Rotation:

The specific rotation of a 0.2% (w/v) aqueous solution of Carrisyn clarified by passing through a .45 μ m membrane filter (Uniflo™, Schleicher & Scheuell Inc., Keene, NH) is:

$$[\alpha]_{589}^{20} = -20 \text{ signifying a B-linkage. Deionized}$$

water was used as the blank.

(d) Alkali Treatment:

Alkali treatment of Carrisyn causes deacetylation which destroys its ability to form a mucilaginous jelly, this indicates that the O-acetylation of Carrisyn influences its viscosity. The product of deacetylation does not dissolve in water apparently due to increased strong hydrogen bonding.

(e) Infrared Spectroscopy:

The functional groups in Carrisyn are indentified by infrared spectroscopy (Fig. 14 & 15). The strong IR absorption bands near 1740 cm^{-1} and 1250 cm^{-1} signify acetylation. Other absorptions located about 3422 cm^{-1} , 1066 cm^{-1} , 1639 cm^{-1} , 897 cm^{-1} , 806 cm^{-1} and 773 cm^{-1} are characteristic of polysaccharides with B-mannosyl linkages. Weak amide I and amide II absorptions are located about 1649 cm^{-1} and 1541 cm^{-1} respectively, (Fig. 14) these are due to protein impurities since when the Carrisyn™ is treated with protease and dialyzed, these peaks are absent (Fig. 15).

(f) Molecular Weight Distribution:

Carrisyn is polydispersed with at least 73% of the material greater than 10,000 daltons as determined by size exclusion chromatography. An example, of a high pressure liquid chromatogram (Figure 13) demonstrates three fractions labeled A, B and C. Peak A represents Carrisyn greater than 100,000 daltons and peak B represents Carrisyn greater than 10,000 daltons but less than 100,000 daltons. Peak C represents the molecular weight constituents of Carrisyn. The area

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of peaks B and C increase as Carrisyn decomposes which in turn causes a decrease in the area of Peak A.

It will be readily apparent to those in the art from the broad teaching of the invention and the illustrative examples that there is the possibility of the substitution of unnamed chemicals and steps for the preferred enumerated chemicals and process steps; the lack of mention of such unnamed chemicals and steps does not indicate that they are not within the scope of the invention; but are omitted only because of the requirement that this teaching be concise and exact.

What is claimed is:

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1. A process for extracting the active chemical substance in the aloe plant from a leaf of the aloe plant, comprising the steps of:

a) obtaining aloe juice having solubilized matter;

b) adjusting the pH of said aloe juice of from about 3.00 to about 3.50;

c) adding a water soluble, lower aliphatic polar solvent to the aloe juice to precipitate the active chemical substance and thereby to form a heterogeneous solution;

d) removing the water soluble, lower aliphatic polar solvent and the solubilized matter from the heterogeneous solution to isolate the precipitated active chemical substance; and

e) drying the precipitated active chemical substance.

2. The process for extracting the active chemical substance in the aloe plant from a leaf of the aloe plant according to claim 1, wherein said aloe juice is obtained by:

a) washing an aloe leaf in a bacteriacidal solution to remove substantially all surface dirt and bacteria;

b) removing at least a first end portion from said washed leaf;

c) draining, preserving and collecting anthraquinone rich sap from said cut and washed leaf;

d) removing rind from said leaf to produce a substantially anthraquinone-free gel fillet; and

e) grinding and homogenizing said substantially anthraquinone-free aloe gel fillet to produce substantially anthraquinone-free aloe juice having solubilized matter.

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3. The process for extracting the active chemical substance in the aloe plant from a leaf of the aloe plant, according to Claim 1, wherein said aloe juice is obtained by:

- a) washing an aloe leaf in a bacteriacidal solution to remove substantially all surface dirt and bacteria;
- b) crushing the washed aloe leaf; and
- c) dialyzing the crushed leaf chemically to remove and separate a substantially anthraquinone-free gel from remaining fractions of the crushed leaf.

4. The process for extracting the active chemical substance in the aloe plant from a leaf of the aloe plant, according to Claim 1, wherein said aloe juice is obtained by:

- a) washing an aloe leaf in a bacteriacidal solution to remove substantially all surface dirt and bacteria;
- b) crushing the washed aloe leaf; and
- c) grinding and homogenizing said crushed aloe leaf to produce aloe juice having solubilized matter.

5. The process for extracting the active chemical substance in the aloe plant from a leaf of the aloe plant, according to Claim 1, wherein said aloe juice is obtained by:

- a) washing an aloe leaf in a bacteriacidal solution substantially to remove surface dirt and bacteria;
- b) grinding said washed aloe leaf;
- c) filtering said ground aloe leaf to remove fibrous material; and
- d) homogenizing said ground aloe leaf to produce an anthraquinone-rich aloe juice having solubilized matter.

6. The process for extracting the active chemical substance in the aloe plant from a leaf of the aloe plant, according to Claim 1, wherein said aloe juice is obtained by:

- a) crushing a leaf of an aloe plant to extrude aloe juice having solubilized matter.

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7. The process for extracting the active chemical substance in the aloe plant from a leaf of the aloe plant, according to Claim 1, wherein said aloe juice is obtained by:

a) washing an aloe leaf in a bacteriacidal solution to remove substantially all surface dirt and bacteria;

b) removing rind from said leaf to produce an aloe gel fillet; and

c) grinding and homogenizing said aloe gel fillet to produce aloe juice having solubilized matter.

8. The process for extracting the active chemical substance in the aloe plant from a leaf of the aloe plant, according to Claim 1, wherein said precipitate is irradiated, whereby said active chemical substance is sterilized and preserved.

9. The process for extracting the active chemical substance in the aloe plant from a leaf of the aloe plant, according to Claim 1, wherein four volumes of said water soluble, lower aliphatic polar solvent are added to one volume of aloe juice to precipitate said active chemical substance.

10. The process for extracting the active chemical substance in the aloe plant from a leaf of the aloe plant, according to Claim 1, wherein said water soluble, lower aliphatic polar solvent is selected from the group consisting of methanol, ethanol and propanol.

11. The process for extracting the active chemical substance in the aloe plant from a leaf of the aloe plant, according to Claim 10, wherein said water soluble, lower aliphatic polar solvent is ethanol.

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12. The process for extracting the active chemical substance in the aloe plant from a leaf of the aloe plant, according to Claim 1, wherein said active chemical substance is precipitated from said water soluble, lower aliphatic polar solvent and aloe juice for approximately four hours before said water soluble, lower aliphatic polar solvent is removed.

13. The process for extracting the active chemical substance in the aloe plant from a leaf of the aloe plant, according to Claim 1, wherein said precipitated active chemical substance is dried by lyophilization.

14. The process for extracting the active chemical substance in the aloe plant from a leaf of the aloe plant, according to Claim 1, further comprising the steps of:

f) redissolving the precipitated active chemical substance in phosphate buffer;

g) treating the redissolved precipitated active chemical substance with non-specific protease followed by extensive dialysis; and

h) lyophilizing the non-filterable product.

15. A product produced by the process of Claim 1.

16. A product produced by the process of Claim 2.

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17. A product produced by the process of Claim 3.
18. A product produced by the process of Claim 4.
19. A product produced by the process of Claim 5.
20. A product produced by the process of Claim 6.
21. A product produced by the process of Claim 7.
22. A product produced by the process of Claim 8.
23. A product produced by the process of Claim 9.
24. A product produced by the process of Claim 10.
25. A product produced by the process of Claim 11.
26. A product produced by the process of Claim 12.
27. A product produced by the process of Claim 13.
28. A product produced by the process of Claim 14.
29. Composition of matter, comprising:
a substantially non-degradable lyophilized polymer of linear β (1 - 4)-D-mannosyl units wherein randomly interspersed acetyl groups are linked to the polymer through an oxygen atom and wherein D-galactopyranose is linked to the polymer through an α (2-6) linkage at a ratio of about one D-galactopyranose residue per seventy monomer units.
30. Composition of matter according to Claim 29 wherein said polymer is irradiated with gamma or microwave radiation.

31. Composition of matter according to Claim 29, wherein the degree of acetylation is about 0.8 acetyl groups per monomer.

$$\begin{array}{c}
 \text{OAc} \\
 | \\
 \cdot \\
 | \\
 \cdot \\
 | \\
 2 \\
 -\text{B}-\underset{\overset{6}{|}}{\text{D-Manp}}-(1-4)-\text{B}-\underset{\overset{3-6}{|}}{\text{D-Manp}}-(1-4)-\text{B}-\underset{\overset{3-6}{|}}{\text{D-Manp}}-(1-4)-\text{B}-\underset{\overset{3-6}{|}}{\text{D-Manp}}- \\
 | \qquad \qquad \qquad | \\
 \cdot \qquad \qquad \qquad \cdot \\
 | \qquad \qquad \qquad | \\
 \cdot \qquad \qquad \qquad \cdot \\
 | \qquad \qquad \qquad | \\
 2 \qquad \qquad \qquad \text{OAc} \\
 -\text{B}-\underset{\overset{2}{|}}{\text{D-galp}}-(1-4)-\text{B}-\underset{\overset{2}{|}}{\text{D-galp}}-
 \end{array}$$

wherein the degree of acetylation is 0.8 acetyl groups per polymer monomer; wherein the galactopyranose units are attached to the polymer at a ratio of approximately one per seventy monomer units; and wherein the mannose to galactose ratio is approximately 20:1.

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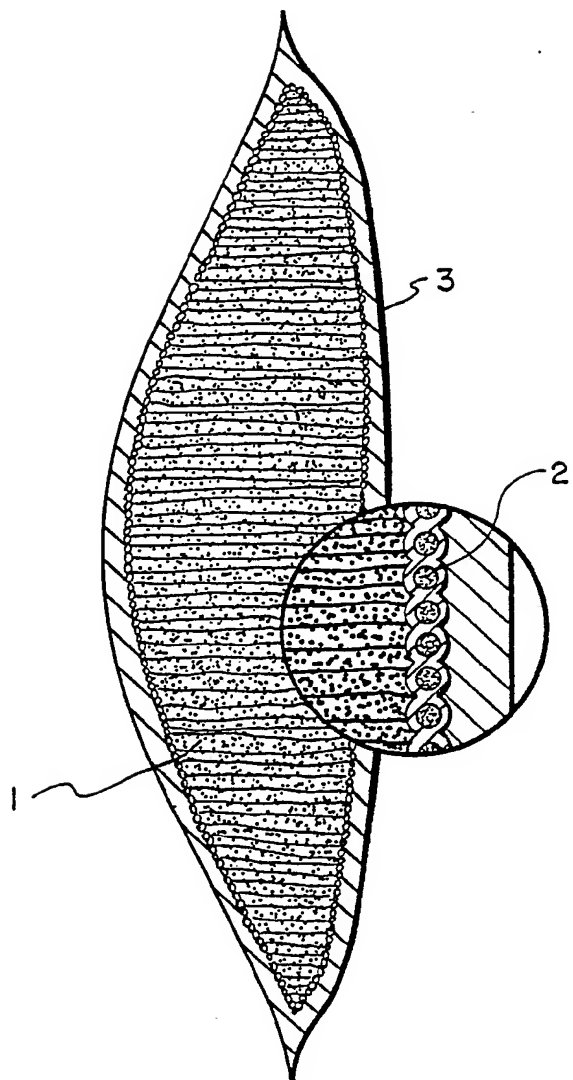


FIG. 1

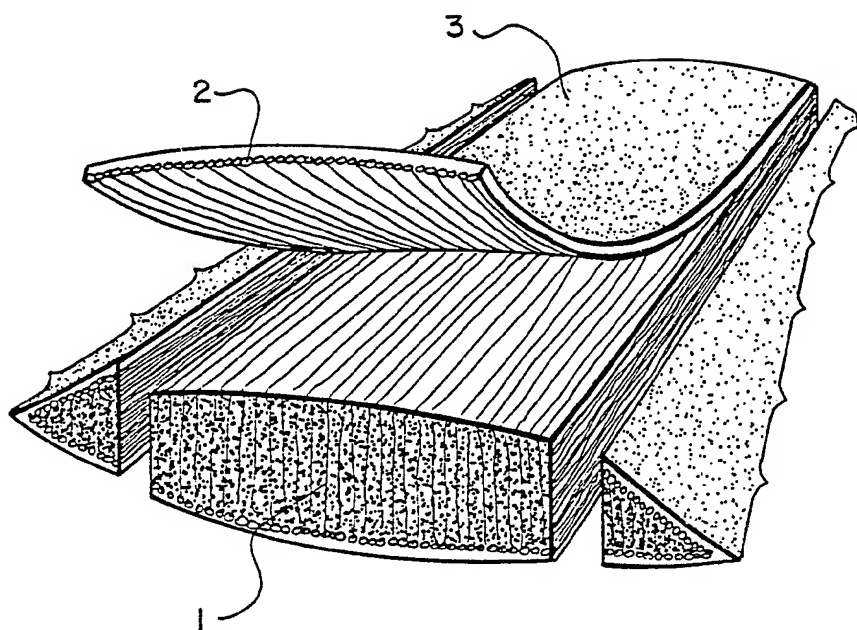


FIG. 2

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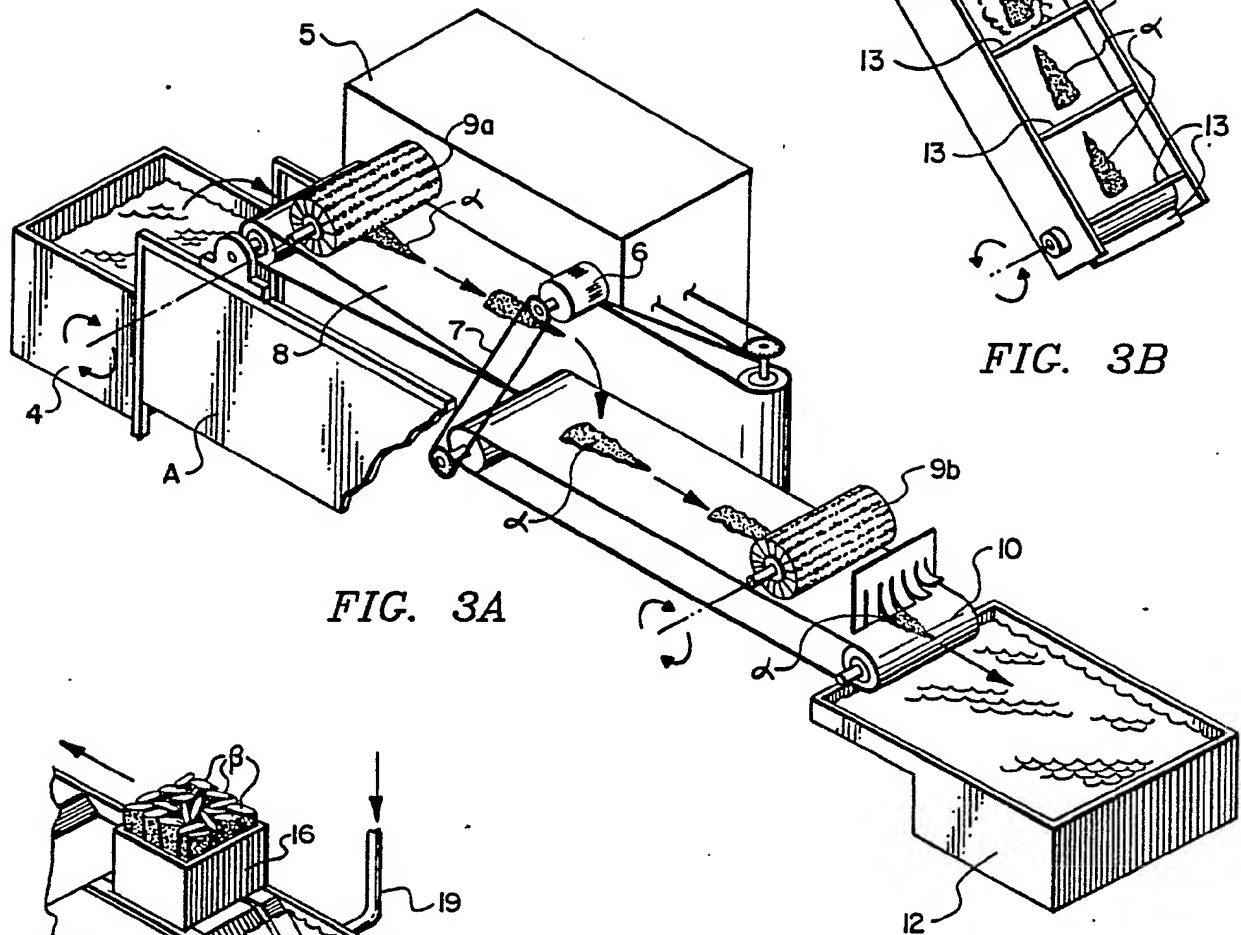


FIG. 3A

FIG. 3B

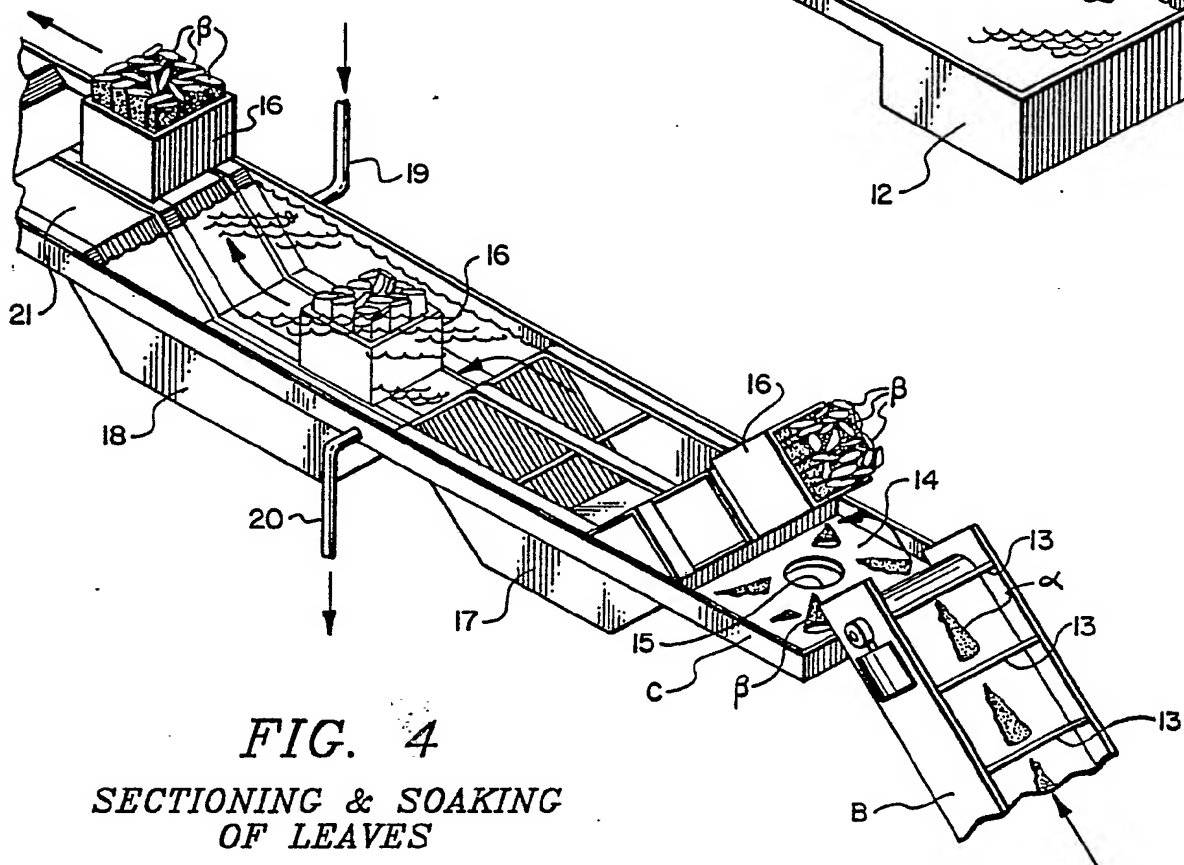


FIG. 4

SECTIONING & SOAKING
OF LEAVES

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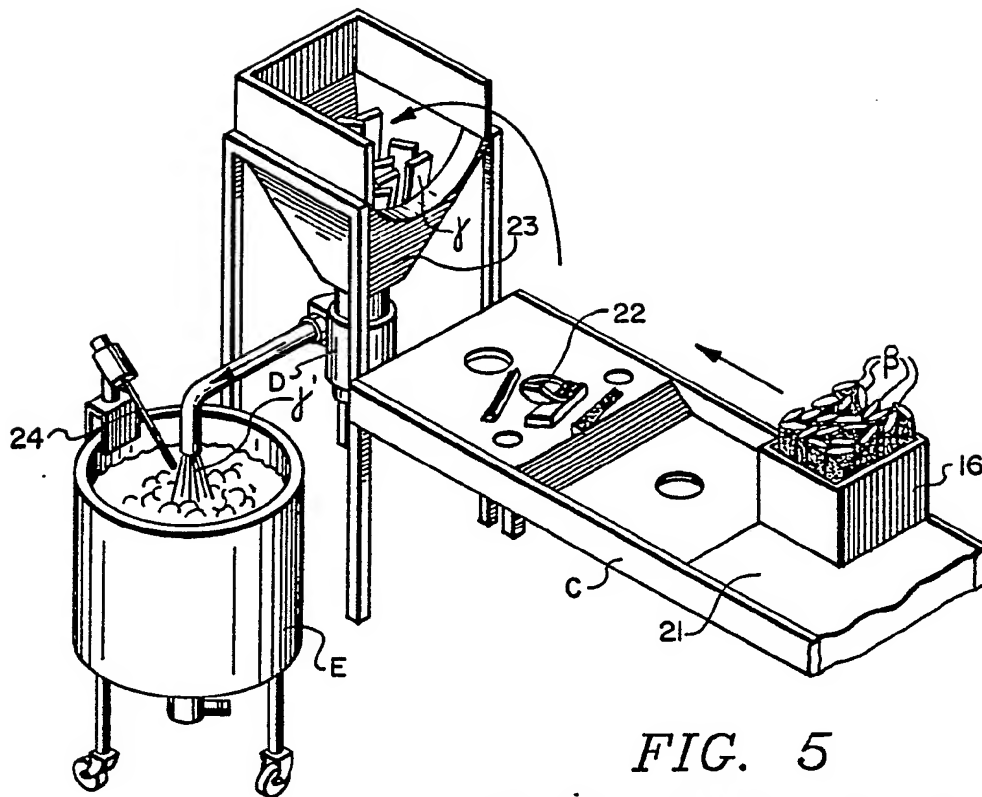
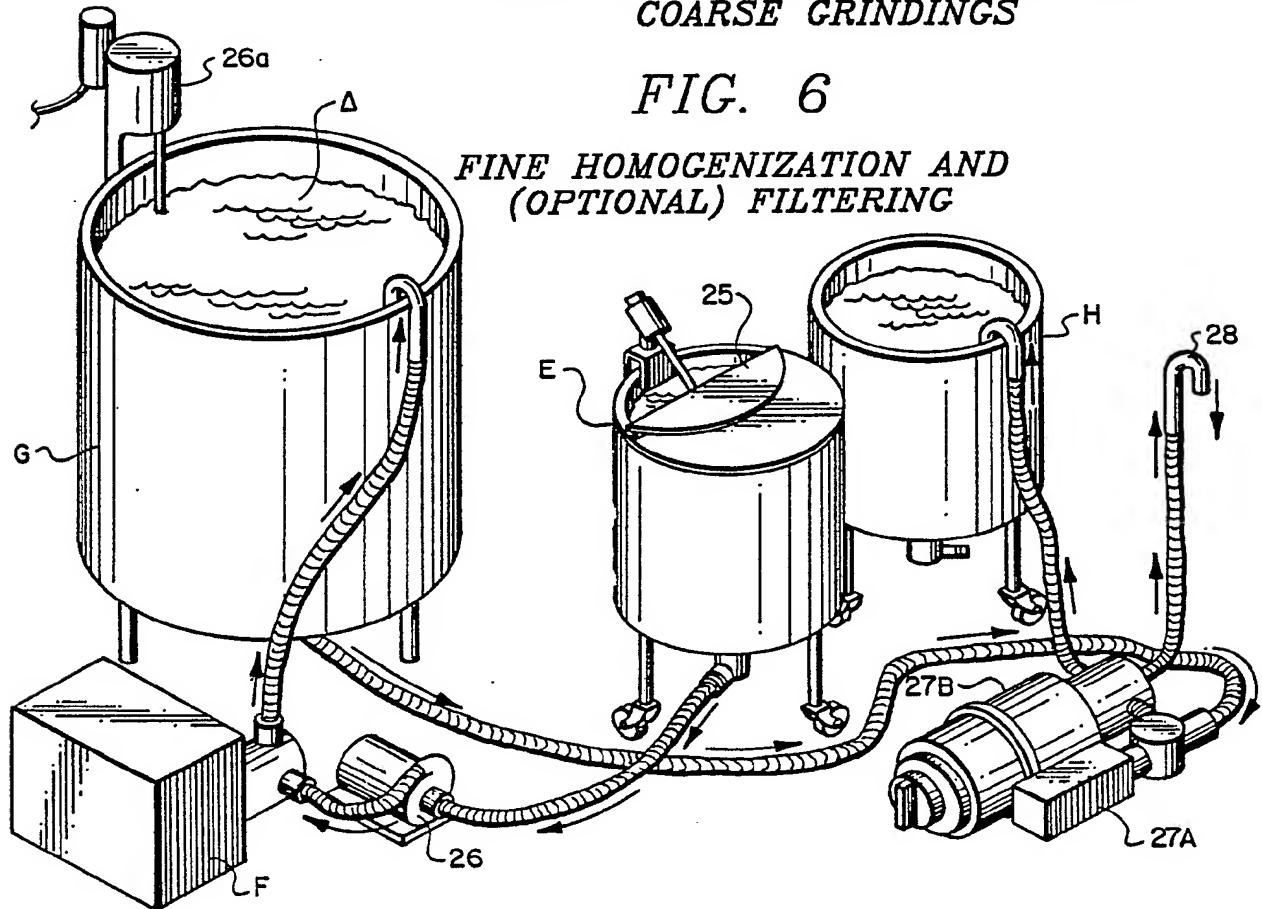


FIG. 5

MANUFACTURING OF FILLETS AND
COARSE GRINDINGS

FIG. 6

FINE HOMOGENIZATION AND
(OPTIONAL) FILTERING

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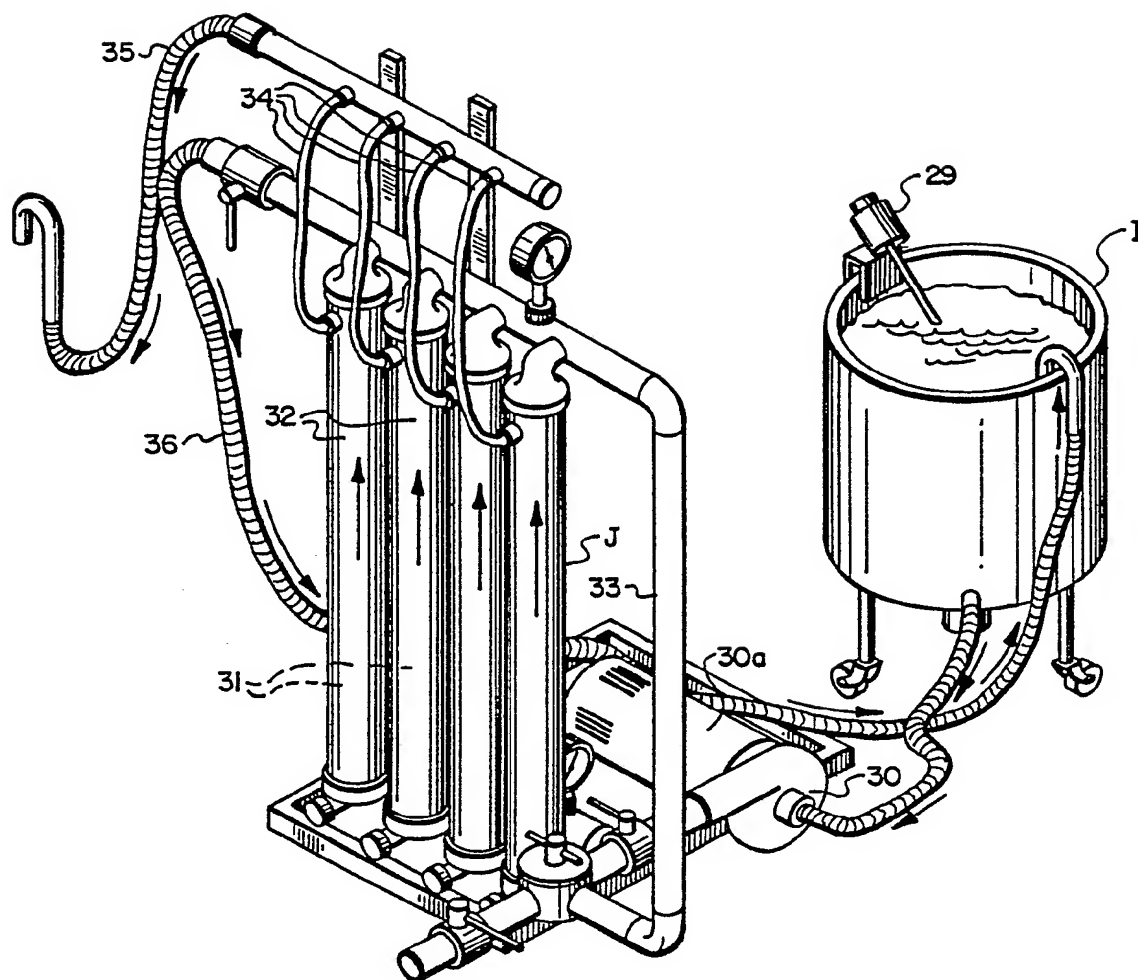


FIG. 7
DIALYSIS UNIT

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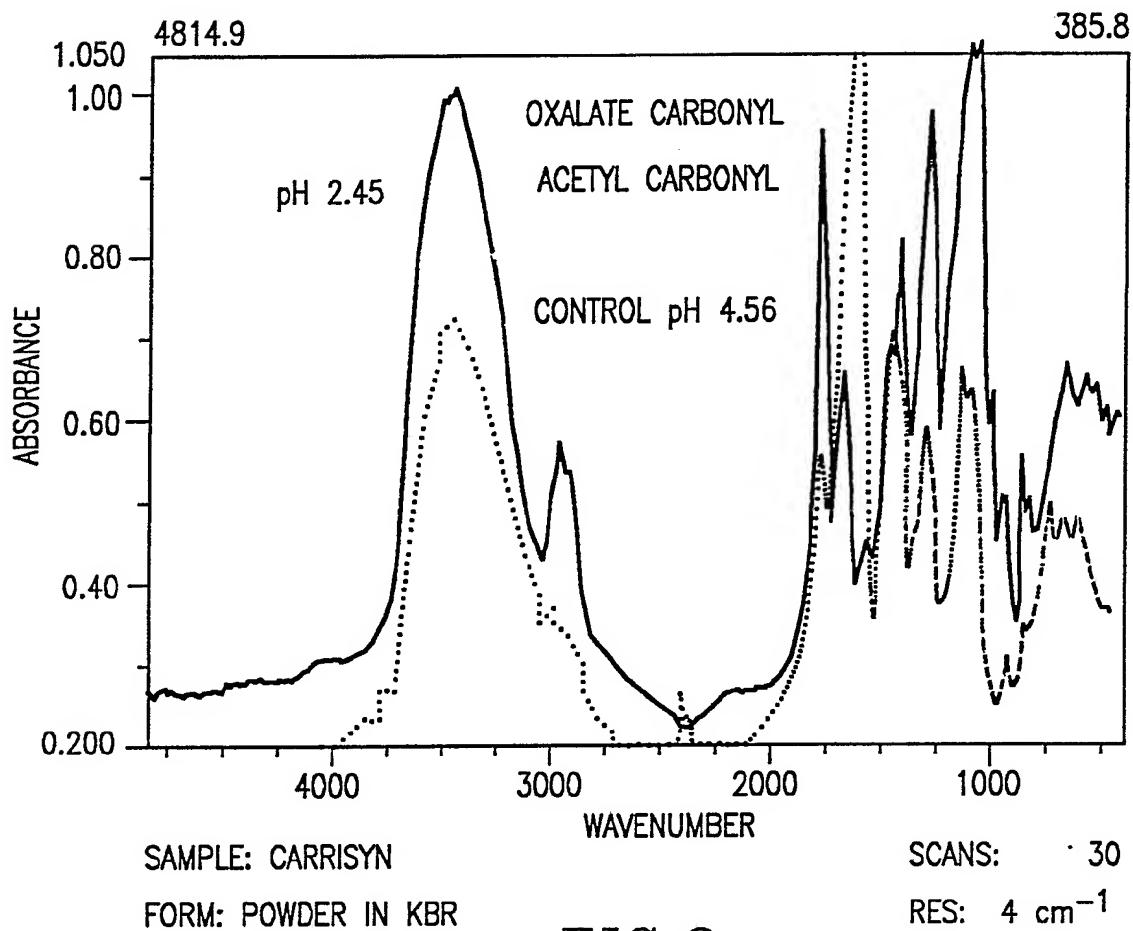


FIG. 8

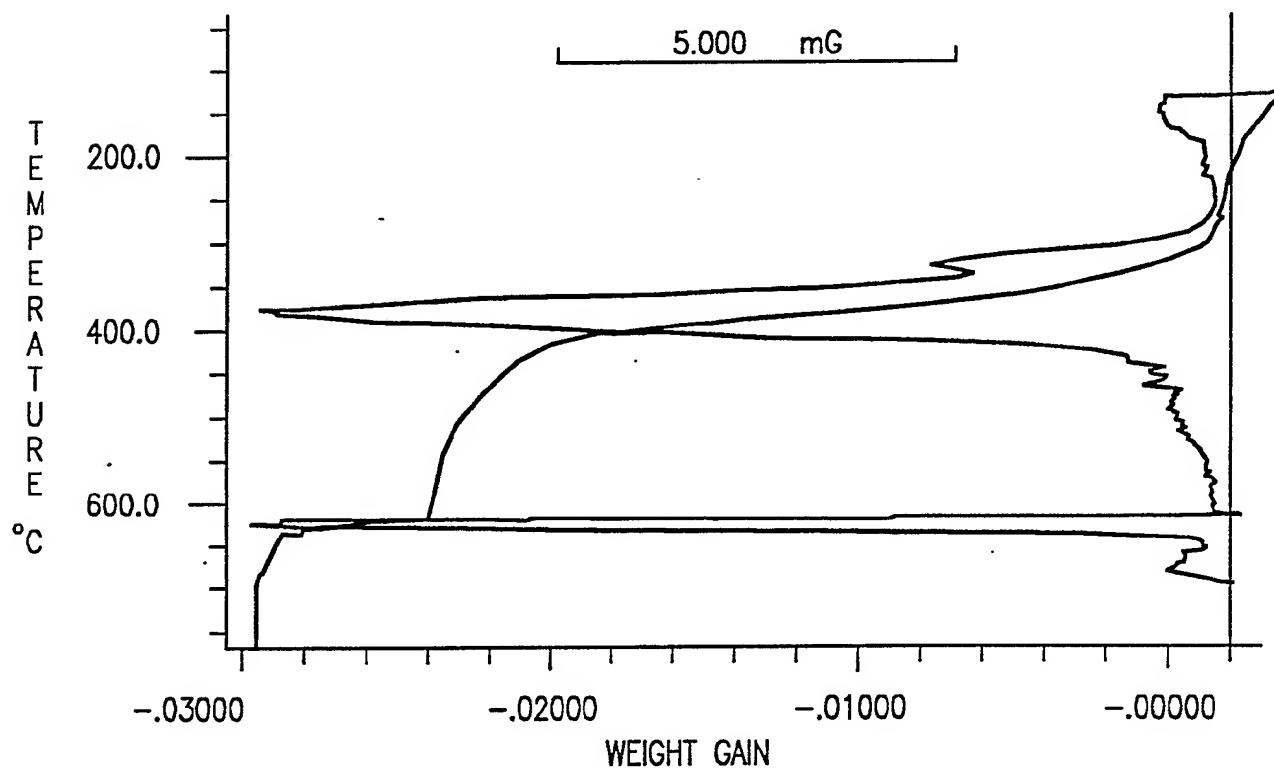
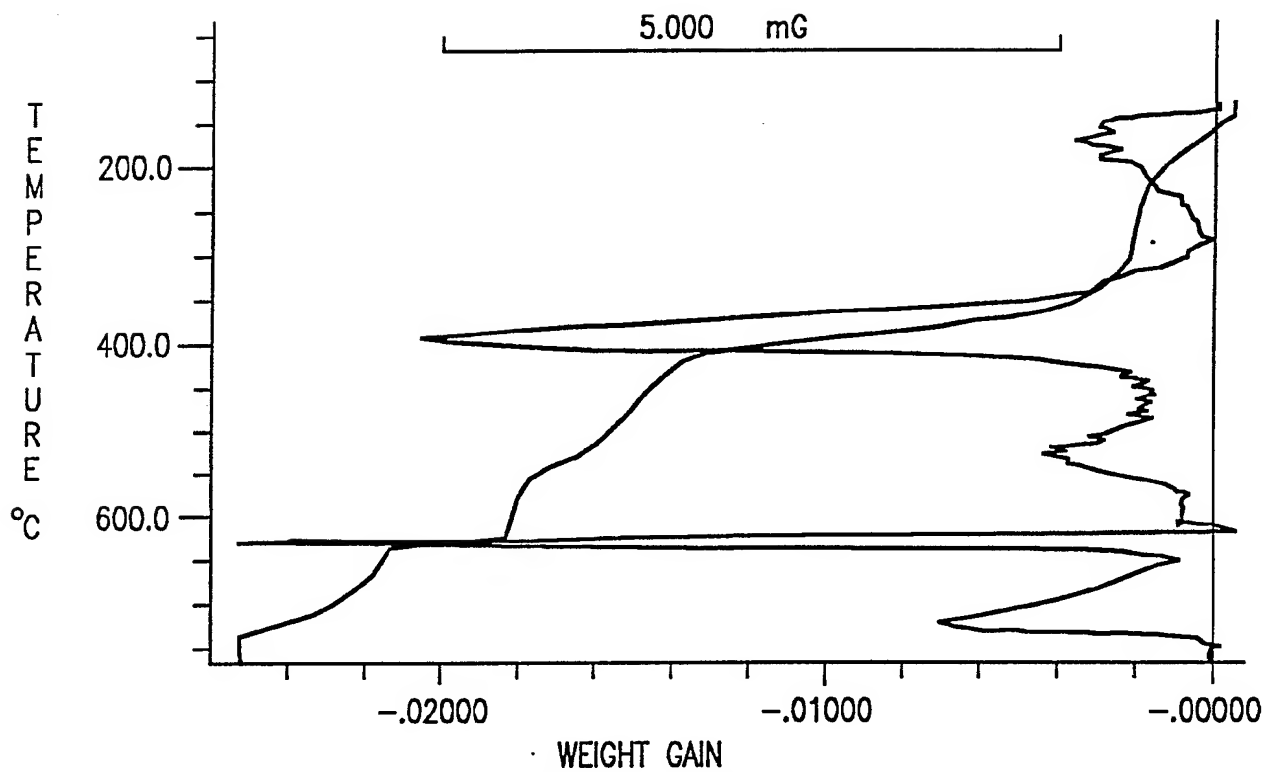
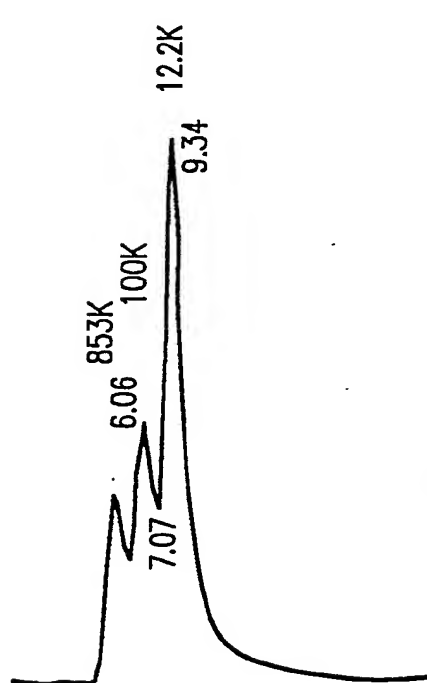
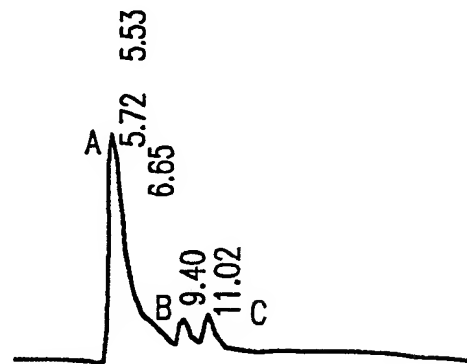


FIG. 9

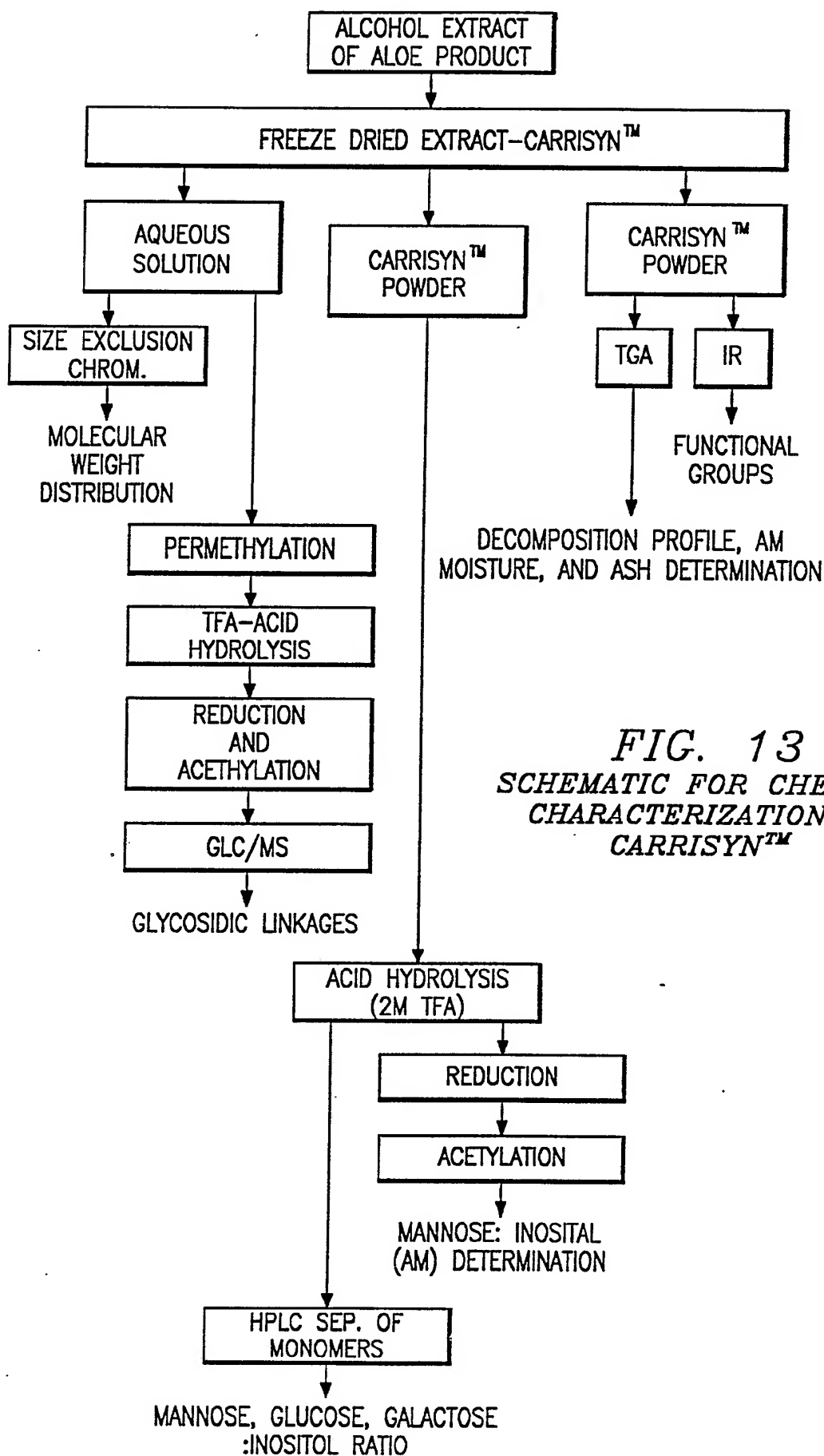
*FIG. 10**FIG. 11*

SIZE EXCLUSION CHROMATOGRAM
OF PULLULAN STANDARDS

*FIG. 12*

SIZE EXCLUSION CHROMATOGRAM
OF CARRISYN

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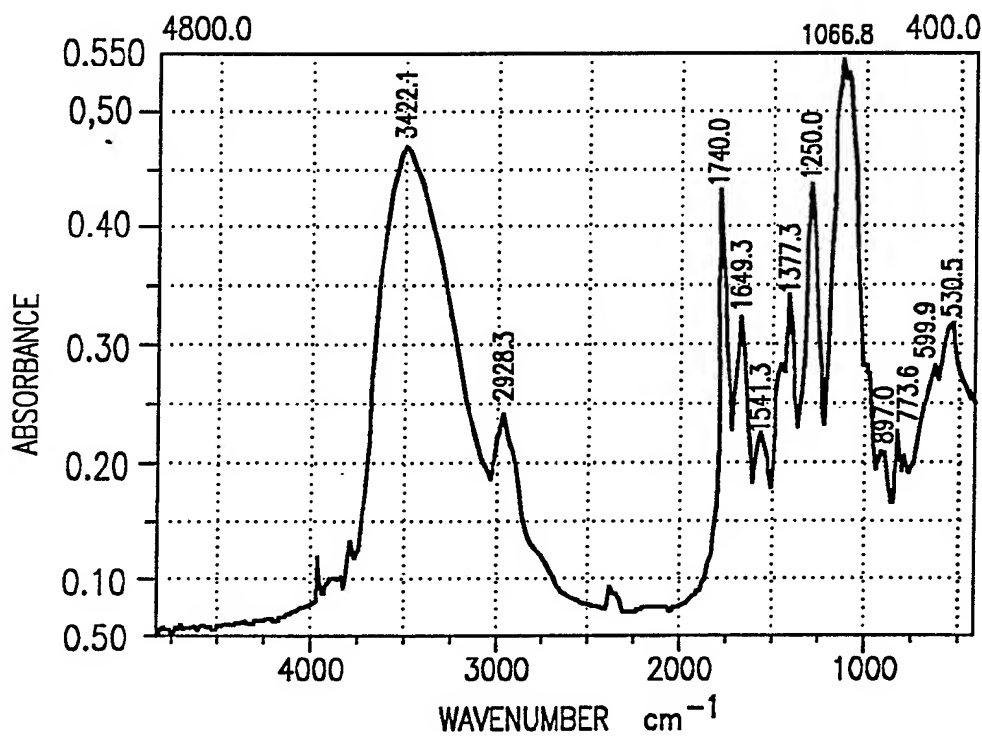


FIG. 14

IR SPECTRUM OF NON-PROTEASE TREATED CARRISYN

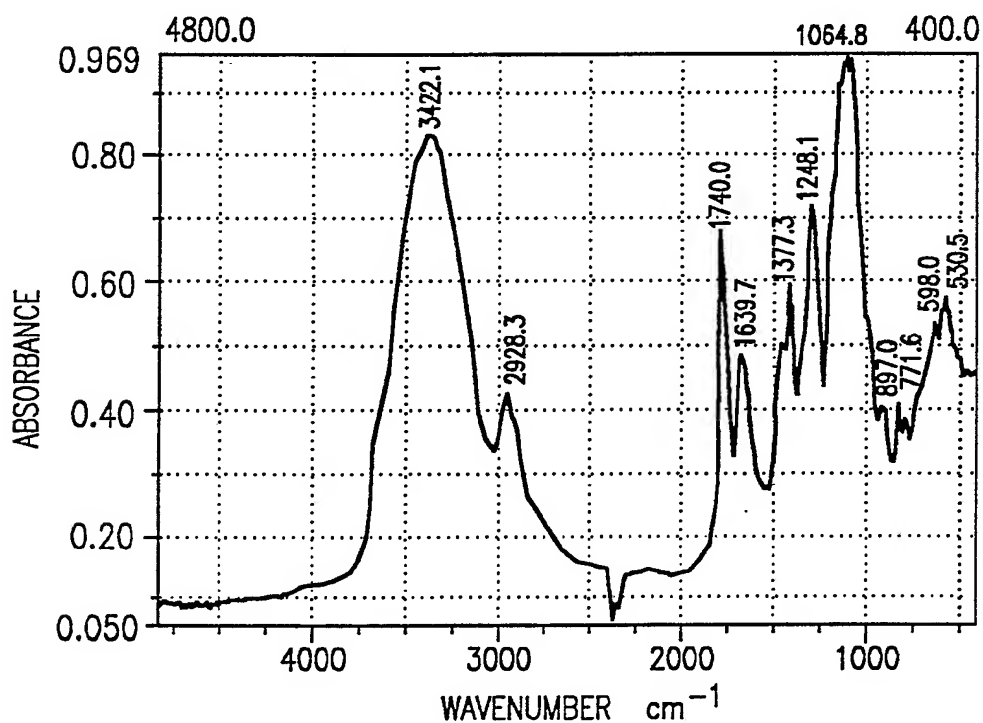


FIG. 15

IR SPECTRUM OF PROTEASE TREATED CARRISYN

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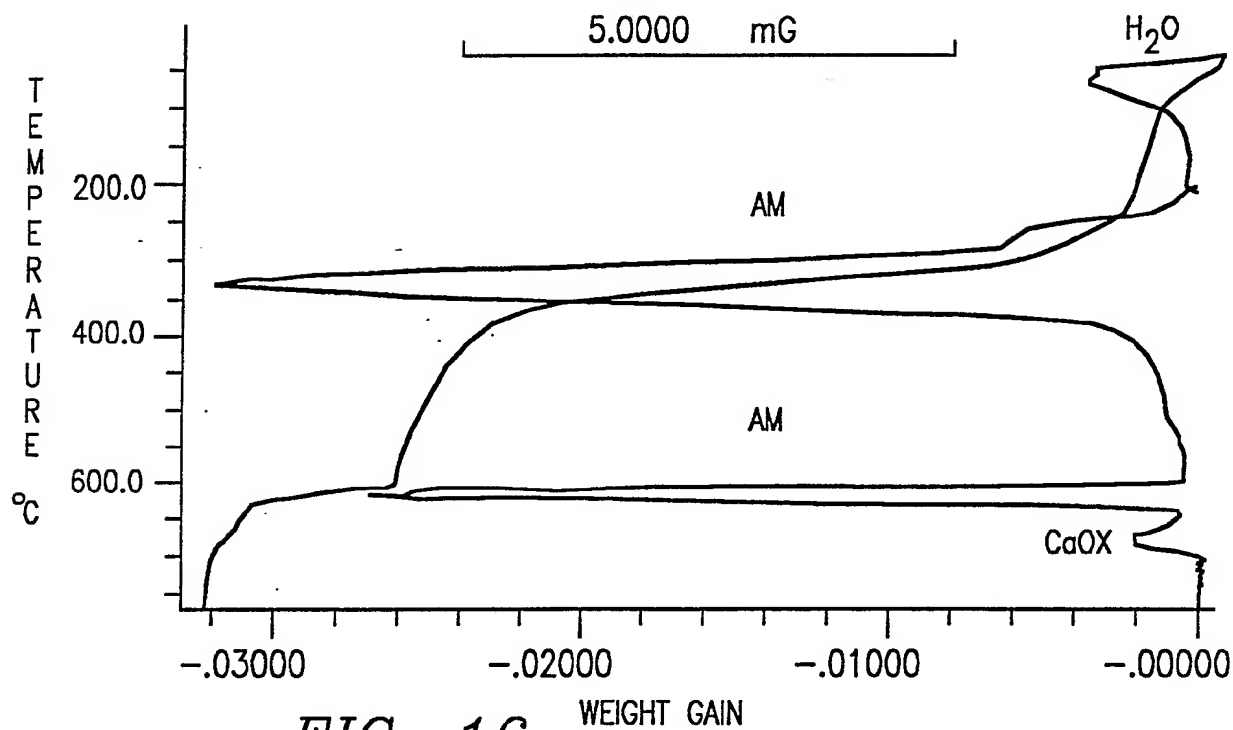


FIG. 16
TGA THERMOGRAM OF CARRISYN

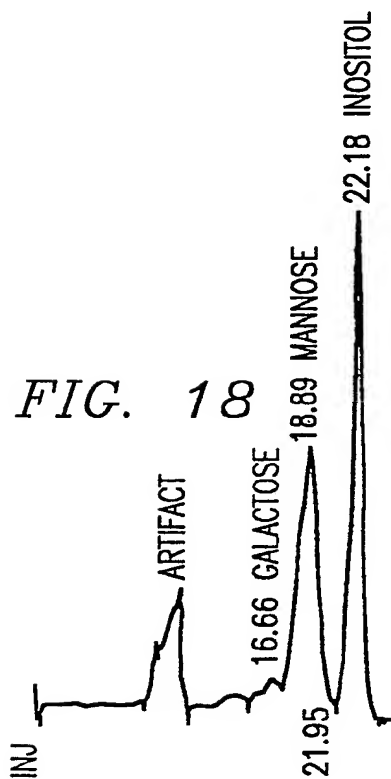


FIG. 18

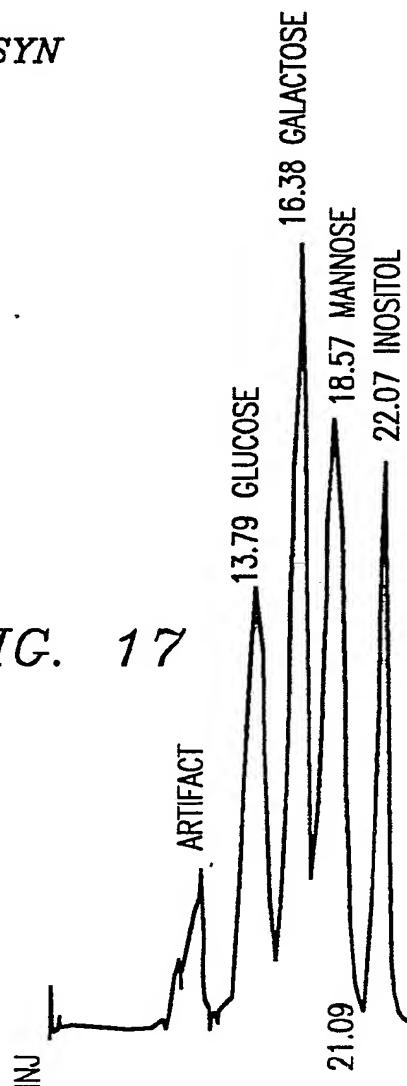


FIG. 17

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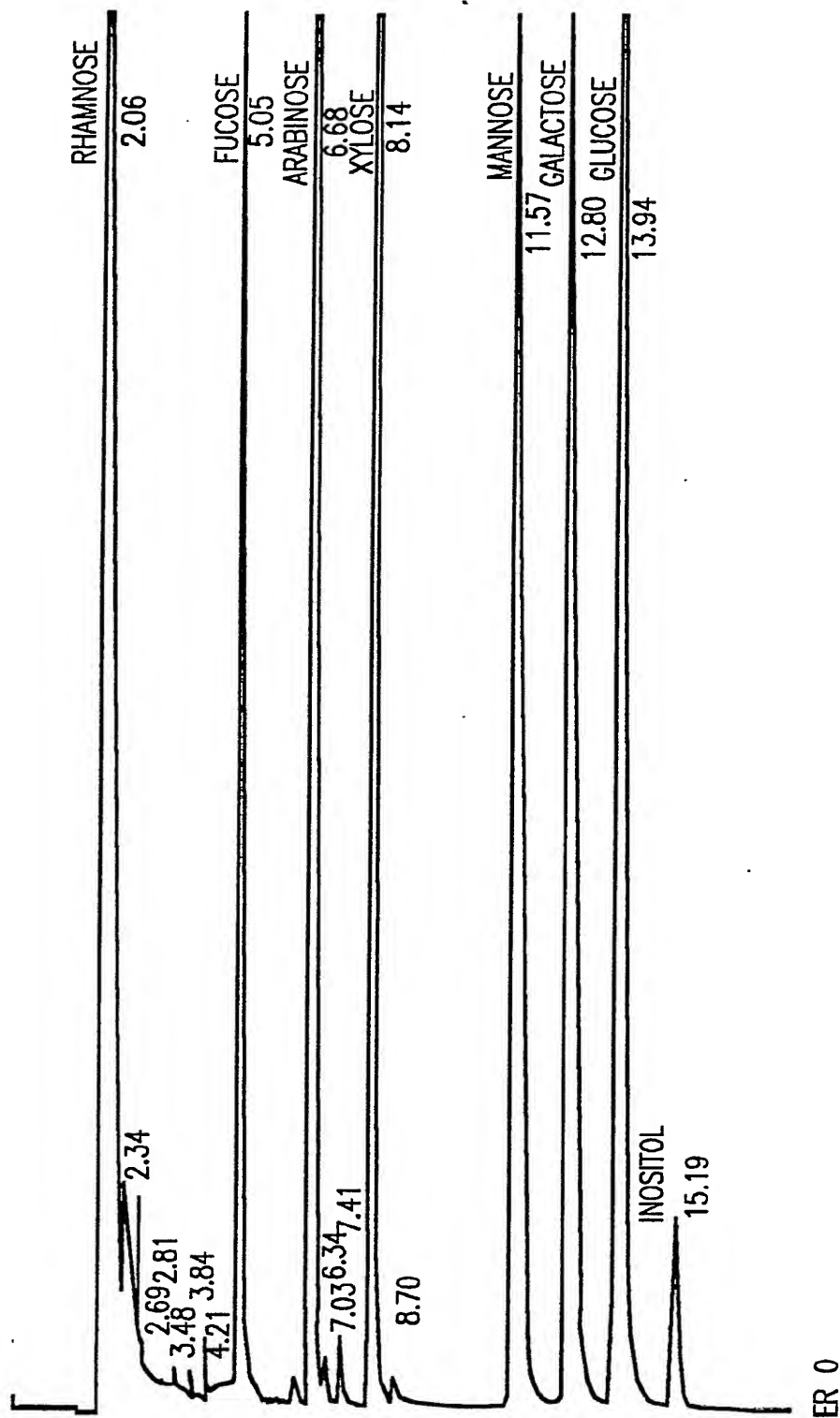


FIG. 19

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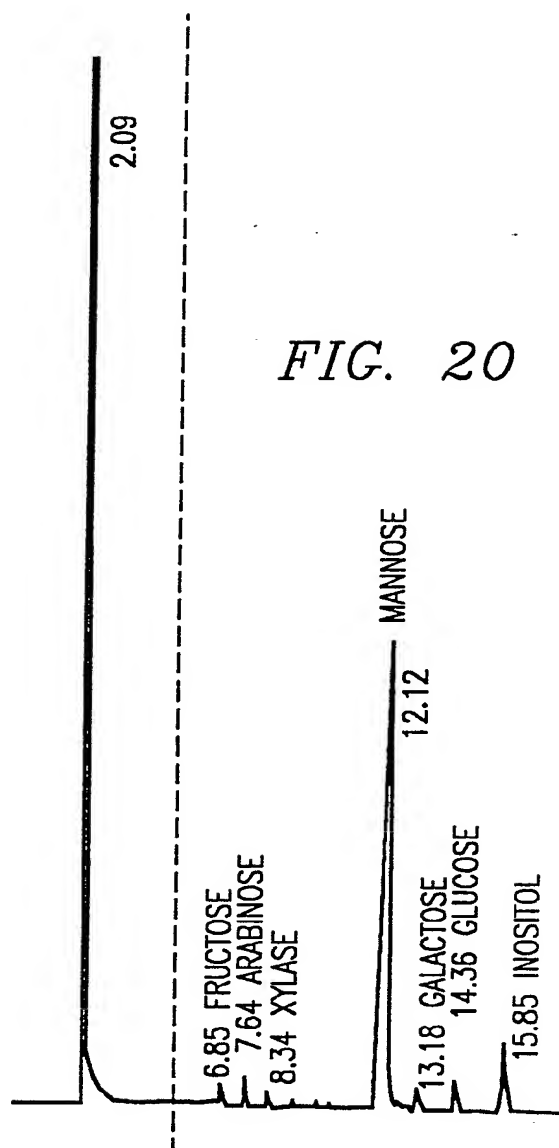


FIG. 20

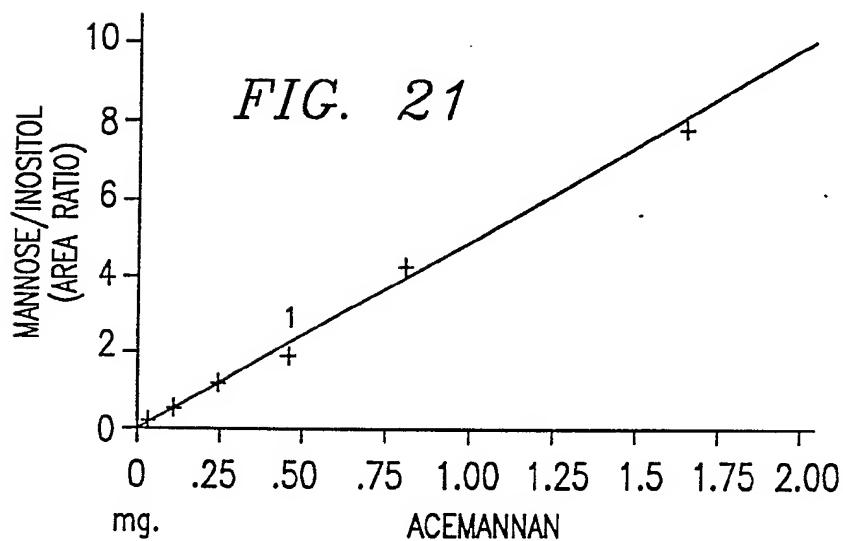
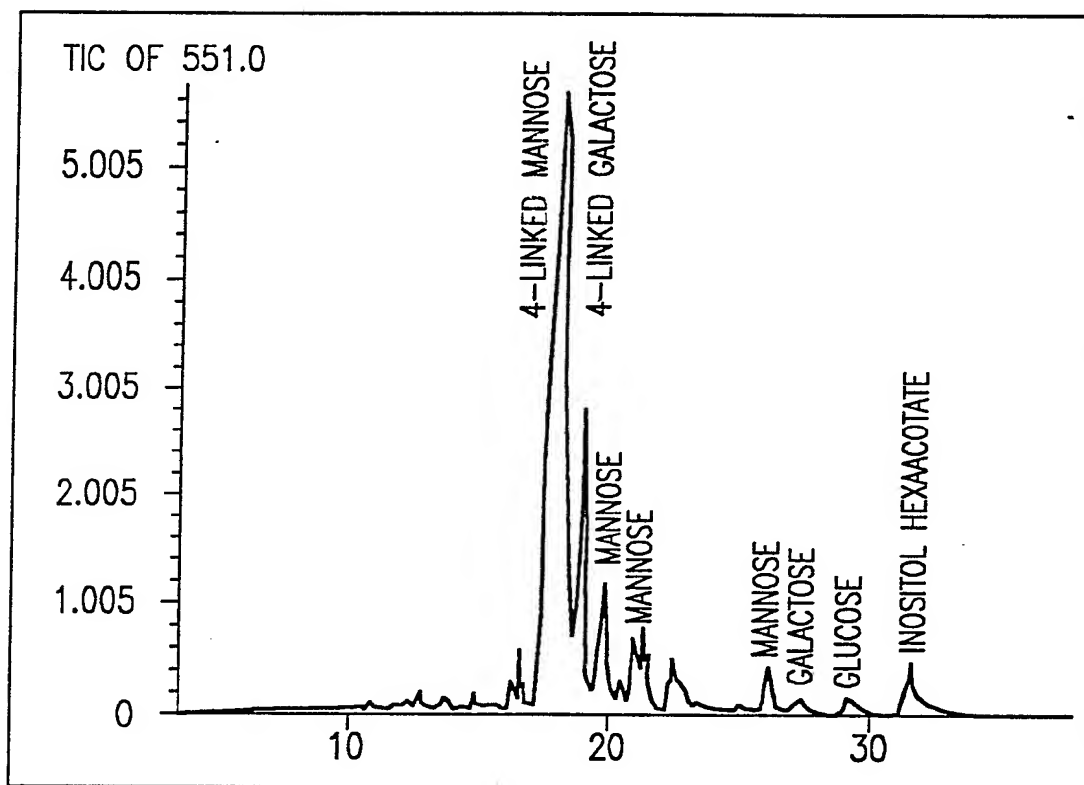


FIG. 21

STANDARD CURVE OF MANNOSE/INOSITOL RATIO
AGAINST AMOUNT OF ACEMANNAN (AM)

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*TOTAL ION CHROMATOGRAM (TIC) OF PARTIALLY
METHYLATED AND PARTIALLY ACETYLATED
GLYCITOL OF CARRISYN*

FIG. 22

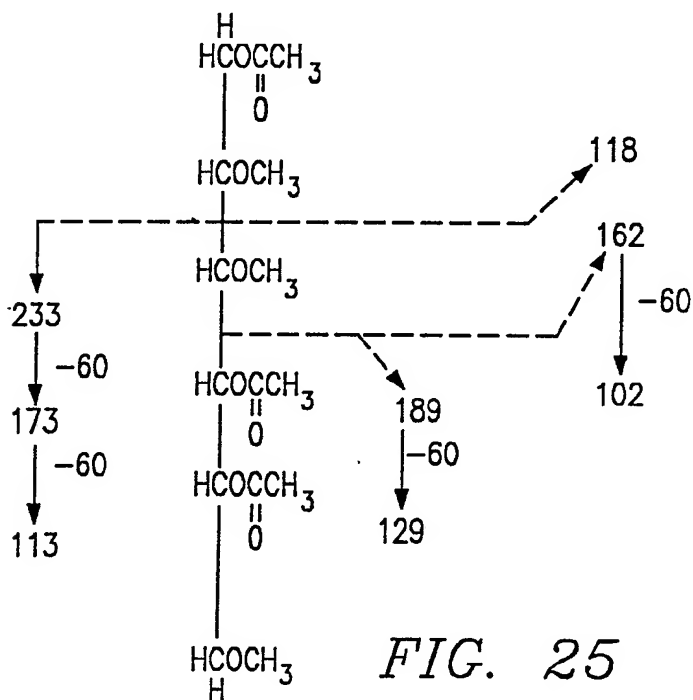


FIG. 25

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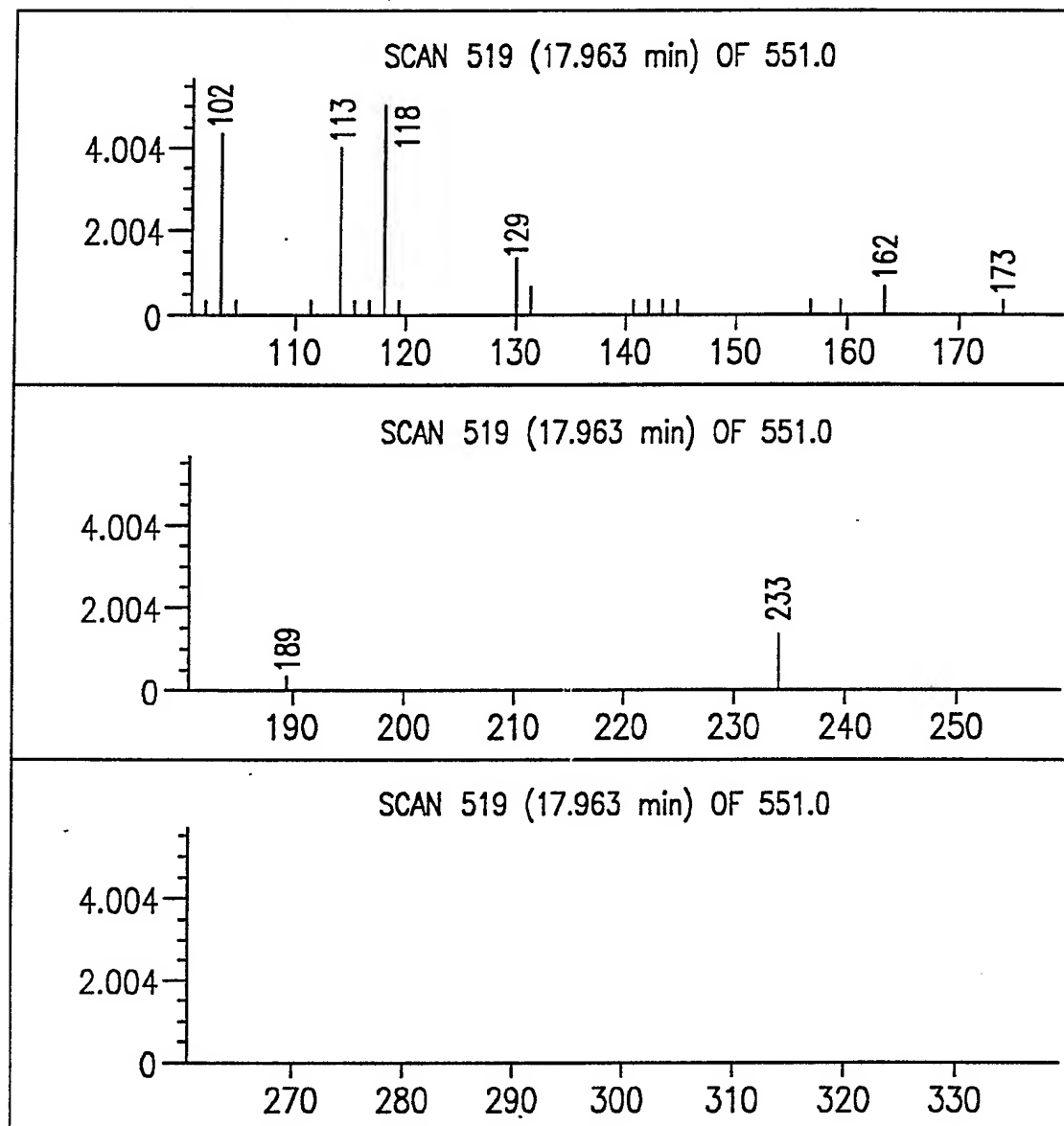


FIG. 23

MASS SPECTRUM OF
1,4,5-TRI-O-ACETYL-2,3,6-TRI-O-METHYL
MANNITOL (4-MANP-).

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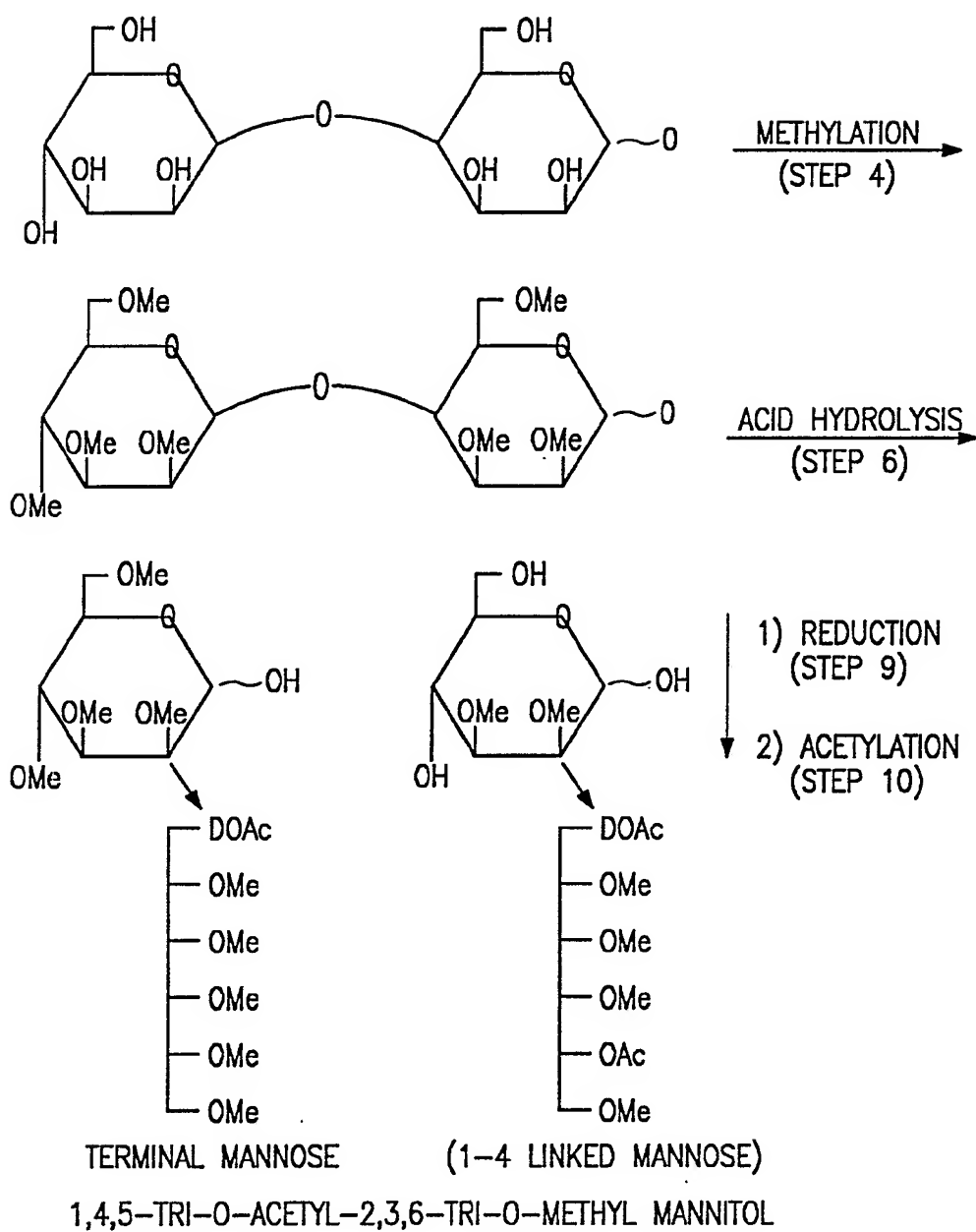



FIG. 24

PARTIALLY METHYLATED MANNITOL ACETATES

INTERNATIONAL SEARCH REPORT

International Application No. PCT/US89/00036

I. CLASSIFICATION OF SUBJECT MATTER (if several classification symbols apply, indicate all) ⁶		
According to International Patent Classification (IPC) or to both National Classification and IPC IPC(4): A61K 35/78, 31/715; C07G 17/00 U.S. CL.: 424/195.1; 514/53, 847, 424/DIG 13; 536/53, 123		
II. FIELDS SEARCHED		
Minimum Documentation Searched ⁷		
Classification System	Classification Symbols	
U.S.	424/195.1; 514/53, 847, 424/DIG 13; 536/53, 123	
Documentation Searched other than Minimum Documentation to the Extent that such Documents are Included in the Fields Searched ⁸		
III. DOCUMENTS CONSIDERED TO BE RELEVANT ⁹		
Category *	Citation of Document, ¹¹ with indication, where appropriate, of the relevant passages ¹²	Relevant to Claim No. ¹³
A	US, A, 3,362,951 FARKAS 09 January 1968.	1-32
A	US, A, 3,892,853 COBBLE 01 July 1975	1-32
A	US, A, 4,555,987 TUMLINSON 03 December 1985	1-32
A	US, A, 3,920,816 SEEGALL, 18 November 1975	1-32
A	US, A, 4,178,372 COATS 11 December 1979	1-32
<div style="display: flex; justify-content: space-between;"> <div style="width: 45%;"> <p>* Special categories of cited documents: ¹⁰</p> <p>"A" document defining the general state of the art which is not considered to be of particular relevance</p> <p>"E" earlier document but published on or after the international filing date</p> <p>"L" document which may throw doubts on priority claim(s) or which is cited to establish the publication date of another citation or other special reason (as specified)</p> <p>"O" document referring to an oral disclosure, use, exhibition or other means</p> <p>"P" document published prior to the international filing date but later than the priority date claimed</p> </div> <div style="width: 45%;"> <p>"T" later document published after the international filing date or priority date and not in conflict with the application but cited to understand the principle or theory underlying the invention</p> <p>"X" document of particular relevance; the claimed invention cannot be considered novel or cannot be considered to involve an inventive step</p> <p>"Y" document of particular relevance; the claimed invention cannot be considered to involve an inventive step when the document is combined with one or more other such documents, such combination being obvious to a person skilled in the art.</p> <p>"Δ" document member of the same patent family</p> </div> </div>		
IV. CERTIFICATION		
Date of the Actual Completion of the International Search	Date of Mailing of this International Search Report	
06 April 1989	11 MAY 1989	
International Searching Authority ISA/US	Signature of Authorized Officer  JOHN W. ROLLINS, JR	